PRACTICE PROBLEMS ON NET IONIC EQUATIONS

Answer Key to Practice Problems on Net Ionic Equations:

Molecular: AgNO₃ (ag) + KCl (ag) → AgCl (s) + KNO₃ (ag)

Total lonic: $Ag^{-}(aq) + NQ_{0}^{-}(aq) + K^{-}(aq) + Cl^{-}(aq) \rightarrow AgCl(s) + K^{-}(aq) + NQ_{0}^{-}(aq)$

Net Ionic: Ag⁻(aq) + Cl (aq) → AgCl(s)

Molecular: Mg(NO₃)₂ (ag) + Na₂CO₂ (ag) → MgCO₃ (s) + 2 NaNO₂ (ag)

Total lonic: Mg2 (aq) + 2-NO₄ (aq) + 2-Na* (aq) + CO₂* (aq) → MgCO₃ (s) + 2-Na* (aq) + 2 NO₄ (aq)

Net lonic: $Mg^{2-}(aq) + CO_2^{-2}(aq) \rightarrow MgCO_3(s)$

Molecular: SrBr₂ (aq) + K₂SO₄ (aq) → SrSO₄ (s) + 2 KBr (aq)

Total lonic: Sr^{3} (ag) + 2-8r (ag) + 2-K* (ag) + SO_{3} (ag) \rightarrow $SrSO_{4}$ (s) + 2-K* (ag) + 2-8r (ag)

Net Ionic: $Sr^{(r)}$ (ag) + $SO_c^{(r)}$ (ag) $\rightarrow SrSO_c$ (s)

Molecular: MnCl₂ (aq) + (NH₄)₂CO₃ (aq) → MnCO₃ (s) + 2 NH₄Cl (aq)

Total Ionic: Mn2- (ag) + 2-GI (ag) + 2-NHu- (ag) + CO2- (ag) → MnCO2 (s) + 2-NHu- (ag) + 2-GI (ag)

Net Ionic: Mn2 (ag) + CO₂2 (ag) → MnCO₂ (s)

Molecular: 2 Cr(NO₃)₃ (aq) + 3 FeSO₄ (aq) → 3 Fe(NO₃)₃ (aq) + Cr₂(SO₄)₃ (aq)

Total lonic; 2 Gr° (eq) + 6 NO_{\circ} (eq) + 3 Fe° (eq) + $3 \text{ SO}_{\circ}^{\circ}$ (eq) $\rightarrow 2 \text{ Gr}^{\circ}$ (eq) + 6 NO_{\circ} (eq) + 3 Fe° (eq) +

3-SQ /- (ag)

Net Ionic: No Reaction

Molecular: K₃PO₄ (aq) + Al(NO₃)₃ (aq) → AlPO₄ (s) + 3 KNO₃ (aq)

Total Ionic: 3 K° (ag) + PO_{4}° (ag) + AP° (ag) + 3 NO_{2} (ag) \rightarrow $AIPO_{4}$ (s) + 3 K° (ag) + 3 NO_{2} (ag)

Net Ionic: AP* (aq) + PO₂3 (aq) → AIPO₂ (s)

Molecular: Bel₂ (aq) + Cu₂SO₄ (aq) → BeSO₄ (aq) + 2 Cul (s)

Total Ionic: $Be^{2r}(aq) + 2I^{-}(aq) + 2Cu^{*}(aq) + SQ_{4}^{2r}(aq) \rightarrow Be^{2r}(aq) + SQ_{4}^{2r}(aq) + 2CuI(s)$

Net Ionic: 2 Cu^{pr} (aq) + $2 \text{ I}^{\text{--}}$ (aq) $\rightarrow 2 \text{ Cul}$ (s)

Molecular: Ni(NO₃)₃ (aq) + 3 KBr (aq) -> NiBr₃ (aq) + 3 KNO₃ (aq)

Total lonic; Ni* (ag) + 3 NO₂ (ag) + 3 K* (ag) + 3 Br (ag) -> Ni* (ag) + 3 NO₂ (ag) + 3 K* (ag) + 3 Br (ag)

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