

EARTH SCIENCE

INTRODUCTION

COMPONENTS OF STUDY

1. **Geology:** Both physical & historical
2. **Meteorology:** Weather & climate
3. **Oceanography:** The oceans
4. **Astronomy:** What is external to the atmosphere

THE FOUR SCHEMES (See The Water Conflict

1. Lithosphere: The solid earth.
2. Hydrosphere: The liquid earth.
3. Atmosphere: The gaseous earth.
4. Biosphere: Where life exists.

CONCERNS IN OUR PHYSICAL ENVIRONMENT

1. Resources: Water, soil, metallic & non-metallic elements, sunlight, wind, etc.
2. Per-capita use of metallic and non-metallic elements is 11 tons per year
3. Renewable vs. non-renewable resources
4. Environmental and population problems

MINERALS

- A mineral is a naturally occurring, inorganic, solid material with a definite atomic structure.
- An ore is a useful metallic mineral.

[illegible]

1. Atoms, in ancient Greek, means "small," i.e., the Ancients believed that atoms to be the smallest particle of matter.
2. Atoms are composed of neutrons, protons and electrons.
 - a. Atomic number: Number of protons in nucleus.
 - b. Atomic weight: Average weight of an atom.
 - c. Isotopes: Forms of an element with identical atomic number but different numbers of neutrons in the nucleus.

III. Properties of Minerals

1. Luster: Appearance or quality of light from surface
2. Color: Nature of light and cause of color
3. Streak: Color of material in powdered form
4. Hardness: Moh's Scale (See Figure below)
5. Scratch Test: Higher numbered materials can scratch lower numbered materials
6. Shape is determined by cleavage, crystal form and fracture

6. **Specific Gravity:**
- The ratio of the weight of a mineral to the weight of an equal volume of water.
 - Density of water = 1 gm/cm³ = 1 gm/ml.
 i.e., Gold = 19.3, aluminium = 2.7, zinc = 7.2

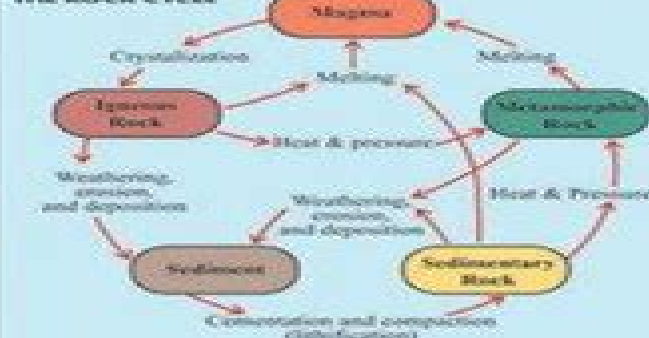
Moh's Scale		Mineral
Hardest	10	Diamond
	9	Corundum
	8	Topaz
	7	Quartz
	6	Orthoclase
	5	Apatite
	4	Fluorite
	3	Calcite
	2	Gypsum
Softest	1	Talc

Figure 1

1. **Silicates:** Minerals with silicic acid and oxygen.
 - a. Silicates in a semi-solid which forms a pyramid-shaped structure with oxygen.
 - b. Other forms include single tetrahedrons (olivine), chains (asbestos, sheets (mica), 3-D structures (feldspar and quartz).
2. **Non-silicates:**
 - a. Make up one-fourth of continental crust
 - b. **Carbonates:** Minerals with carbon and oxygen including calcite, from which we process limestone (roads) and marble (decorative slabs)
 - c. **Oxides:** Oxygen based solids, i.e. iron
 - d. **Sulfides (S):** sulfides (FeS), **Sulfates (SO₄),** **Halides (Cl, F)**
 - e. **Native:** Natural form of self
 - f. **Elemental:** carbon, cobalt
 - g. **Native metals:** iron, zinc, gold without rock

ROCK TYPES

THE RIVER CYCLE



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1. **Extrusive** igneous rocks are formed in volcanic eruptions at surface.
2. **Intrusive** (or **plutonic**) igneous rocks are produced within the earth.
3. **Magma**
 - a. Consists mainly of silicate materials.
 - b. Has gases, such as water vapor.
 - c. Originates from as deep as 290 km.
 - d. Differs in rate of cooling, chemical composition, and amount of gases.
 - e. Produces intrusive (plutonic) igneous rocks.
4. **Lava**
 - a. Similar to magma.
 - b. Most gaseous elements have escaped.
 - c. Produces extrusive (volcanic) igneous rocks.

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- 1. **Crystallization** (ordered pattern of ions)
 - a. the rate of cooling strongly affects crystal size
 - b. slow cooling = large crystals
 - c. quick cooling = small crystals
- 2. **Amorphous** (unordered pattern of ions)
 - a. occurs with instant cooling (quenching)
 - b. produces rocks referred to as glass

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1. By texture, we use arrangement of interlocking crystals.
 - a. **fine-grained**
 - i. forms at surface or within upper crust
 - ii. openings left by gas bubbles are called vesicles
 - iii. common shapes include: **tridymite**
 - iv. **porphyritic** - large crystals embedded in matrix of smaller crystals
2. By mineral composition
 - a. Depends on composition of magma and cooling
 - b. Some magma can create rocks of varying mineral content.
 - c. As magma cools, certain minerals crystallize first (e.g. basalt, granite)

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2. **Metamorphism:** To change form.
3. **Regional metamorphism:** Material under intense stresses, high temperatures.
4. **Contact (thermal) metamorphism:** Changes caused by proximity to magma.
5. **Low grade metamorphism:** Shale becomes slate.
6. **High grade metamorphism:** Gneiss becomes rock.
7. **Often forms during mountain building.**
8. **Some metamorphic rocks are harder than sedimentary rocks.**
9. **Metamorphism changes texture.**
10. **Low-grade metamorphism makes material more compact, more dense.**
11. **Foliated texture:** Particles of material are brought into line with one another.

III. Aggregates and Well-Being

1. **Heat**
 - a. most important agent
 - b. provides energy for chemical reactions
 - c. For example: clay pottery shatters into a million at great temperature
2. **Pressure**
3. **Chemical Activity:** Most common chemical agent in water

C. Types of Metamorphic Rocks

1. **Hornfels:** Fine grained, dark bluish rock with randomly arranged minerals
2. **Slate:** Fine grained, often gray, foliated rock, split easily along cleavage, planes of weak weakness aligned by pressure
3. **Mylonite:** Dark, fine grained, foliated rock, formed from brecciated or irregularly fractured rock
4. **Phyllite:** Silty, foliated rock, more coarsely grained than slate
5. **Schist:** Foliated rock, more coarsely grained and of higher metamorphic grade than Phyllite
6. **Gneiss:** Foliated, banded rock of the highest metamorphic grade
7. **Quartzite:** Very hard, granular rock, formed from sandstone

RELATIONSHIPS BETWEEN THE TWO GROUPS

If a numeral, which has already formed, occurs in the magma, it will react with the remaining magma to produce the next numeral in the sequence. For example, silver forms first; silver then reacts with remaining magma to form mercury.

18. Identifying Similar and Other Dissimilar Numbers

1. Calcium-rich parts of the magmas form small crystals of feldspar.
2. Their react with sodium in the magma to form more and more sodium-rich.

Fig. 1. Model of the system.

1. When everything is almost cool, combine the sugar with the high altitude corn meal, and pour it all over.
2. When cooking is complete, everything that cooked in the same time will be done by the time the corn is cooked (different times and quantities could mean one another's ruin).



High School Earth Science Proficiency Study Guide

David Kirk



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