

Solutions Manual for Fluid Mechanics 5th Edition by Kundu Cohen and Dowling

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Chapter 1 Introduction

1.1 A gas at 20°C may be *rarefied* if it contains less than 10^{12} molecules per mm^3 . If Avogadro's number is $6.023\text{E}23$ molecules per mole, what air pressure does this represent?

Solution: The mass of one molecule of air may be computed as

$$m = \frac{\text{Molecular weight}}{\text{Avogadro's number}} = \frac{28.97 \text{ mol}^{-1}}{6.023\text{E}23 \text{ molecules/g} \cdot \text{mol}} = 4.81\text{E}-23 \text{ g}$$

Then the density of air containing 10^{12} molecules per mm^3 is, in SI units,

$$\begin{aligned} \rho &= \frac{(10^{12} \text{ molecules})}{\text{mm}^3} \cdot \frac{4.81\text{E}-23 \text{ g}}{\text{molecule}} \\ &= 4.81\text{E}-11 \frac{\text{g}}{\text{mm}^3} = 4.81\text{E}-5 \frac{\text{kg}}{\text{m}^3} \end{aligned}$$

Finally, from the perfect gas law, Eq. (1.13), at 20°C = 293 K, we obtain the pressure:

$$p = \rho RT = 4.81\text{E}-5 \frac{\text{kg}}{\text{m}^3} \cdot \frac{\text{m}^2}{\text{s}^2 \cdot \text{K}} (293 \text{ K}) = 4.0 \text{ Pa} \quad \text{Ans.}$$

1.2 The earth's atmosphere can be modeled as a uniform layer of air of thickness 20 km and average density 0.6 kg/m^3 (see Table A-6). Use these values to estimate the total mass and total number of molecules of air in the entire atmosphere of the earth.

Solution: Let R_e be the earth's radius $\approx 6377 \text{ km}$. Then the total mass of air in the atmosphere is

$$\begin{aligned} m_a &= \rho \, d\text{Vol} = \rho_{\text{avg}} (\text{Air Vol}) \approx \rho_{\text{avg}} 4\pi R_e^2 (\text{Air thickness}) \\ &= (0.6 \text{ kg/m}^3) 4\pi (6.377\text{E}6 \text{ m})^2 (20\text{E}3 \text{ m}) \approx 6.1\text{E}18 \text{ kg} \quad \text{Ans.} \end{aligned}$$

Dividing by the mass of one molecule $\approx 4.8\text{E}-23 \text{ g}$ (see Prob. 1.1 above), we obtain the total number of molecules in the earth's atmosphere:

$$N_{\text{molecules}} = \frac{m(\text{atmosphere})}{m(\text{one molecule})} = \frac{6.1\text{E}21 \text{ grams}}{4.8\text{E}-23 \text{ gm/molecule}} \approx 1.3\text{E}44 \text{ molecules} \quad \text{Ans.}$$

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Fluid Mechanics Kundu Solutions Manual:

The book delves into Fluid Mechanics Kundu Solutions Manual. Fluid Mechanics Kundu Solutions Manual is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Fluid Mechanics Kundu Solutions Manual, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 4: Fluid Mechanics Kundu Solutions Manual in Specific Contexts
 - Chapter 5: Conclusion
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Table of Contents Fluid Mechanics Kundu Solutions Manual

1. Understanding the eBook Fluid Mechanics Kundu Solutions Manual
 - The Rise of Digital Reading Fluid Mechanics Kundu Solutions Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Fluid Mechanics Kundu Solutions Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Fluid Mechanics Kundu Solutions Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Fluid Mechanics Kundu Solutions Manual
 - Personalized Recommendations
 - Fluid Mechanics Kundu Solutions Manual User Reviews and Ratings
 - Fluid Mechanics Kundu Solutions Manual and Bestseller Lists
5. Accessing Fluid Mechanics Kundu Solutions Manual Free and Paid eBooks
 - Fluid Mechanics Kundu Solutions Manual Public Domain eBooks
 - Fluid Mechanics Kundu Solutions Manual eBook Subscription Services
 - Fluid Mechanics Kundu Solutions Manual Budget-Friendly Options
6. Navigating Fluid Mechanics Kundu Solutions Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Fluid Mechanics Kundu Solutions Manual Compatibility with Devices
 - Fluid Mechanics Kundu Solutions Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Fluid Mechanics Kundu Solutions Manual
 - Highlighting and Note-Taking Fluid Mechanics Kundu Solutions Manual
 - Interactive Elements Fluid Mechanics Kundu Solutions Manual

8. Staying Engaged with Fluid Mechanics Kundu Solutions Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Fluid Mechanics Kundu Solutions Manual
9. Balancing eBooks and Physical Books Fluid Mechanics Kundu Solutions Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Fluid Mechanics Kundu Solutions Manual
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Fluid Mechanics Kundu Solutions Manual
 - Setting Reading Goals Fluid Mechanics Kundu Solutions Manual
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Fluid Mechanics Kundu Solutions Manual
 - Fact-Checking eBook Content of Fluid Mechanics Kundu Solutions Manual
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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