

4.7 CHEMISTRY (13%)

4.7.1 Chemistry Paper 1 (13%)

4. (a) Ether (1 mark)
- (b) Add water to each compound.
Add Na_2CO_3 or NaHCO_3 to a suitable part of each sample.
 $\text{C}_2\text{H}_5\text{CO}$ no effervescence (1 mark)
 $\text{C}_2\text{H}_5\text{CO}$ effervescence (1 mark)
on add acidified potassium dichromate (VII) and warm
 $\text{C}_2\text{H}_5\text{CO}$ colour from orange to green
 $\text{C}_2\text{H}_5\text{CO}$ no change
on add acidified Potassium Manganate (VII)
 $\text{C}_2\text{H}_5\text{CO}$ no decolourised
 $\text{C}_2\text{H}_5\text{CO}$ no change (1 mark)
5. (a) Ether (Na_2CO_3) (1 mark)
- (b) — Sodium is very reactive (more electropositive) (1 mark)
— More reactive than carbon
6. (a) Ether
Sodium lamp, used in molecular reactions
Sodium vapour, sodium amalgams
(Na_2CO_3 , Electrolysis and chlorides, etc.) (1 mark)
7. (a) Endothermic change, when one mole of crystal lattice is broken into its ions in gaseous state. (1 mark)
- (b) Endothermic reaction (percent) (1 mark)
8. (a) Boyle's law: The volume of a fixed mass of a gas is inversely proportional to its pressure or the product of pressure and volume is constant at a fixed temperature. (1 mark)

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Iseb Iseb



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