

4.7 CHEMISTRY (13%)

4.7.1 Chemistry Paper 1 (13%)

1. (a) Ether (1 mark)
- (b) Add water to each compound.
Add Na_2CO_3 or NaHCO_3 to a suitable part of each sample.
 $\text{C}_2\text{H}_5\text{CO}$ no effervescence (1 mark)
 $\text{C}_2\text{H}_5\text{CO}_2$ effervescence (1 mark)
no add. solidified/purposive effervescence (VT) and water
 $\text{C}_2\text{H}_5\text{CO}_2$ more likely change to green
 $\text{C}_2\text{H}_5\text{CO}_2$ no change
no add. solidified (Potassium Manganate (VII))
 $\text{C}_2\text{H}_5\text{CO}_2$ will be decolourised
 $\text{C}_2\text{H}_5\text{CO}_2$ no change
2. (a) Ether (Na_2CO_3) (1 mark)
- (b) — Sodium is very reactive (more electropositive) (1 mark)
— More reactive than carbon
3. (a) Ether
Sodium lamp, used in nuclear reactions
Sodium vapour, sodium amalgams
(Na_2CO_3 , Electrodes and cathodes, etc.) (1 mark)
4. (a) Endothermic change, which one order of crystal lattice is broken into its ions in gaseous state. (1 mark)
- (b) Endothermic reaction (percent) (1 mark)
5. (a) Boyle's law: The volume of a fixed mass of a gas is inversely proportional to its pressure or the product of pressure and volume is constant at a fixed temperature. (1 mark)

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Thomas Bond, Chris Hughes



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