

CHEMISTRY



Periodic Table of Elements

1 H 1.008																	18 Ar 39.948
2 He 4.0026																	36 Kr 83.80
3 Li 6.941	4 Be 9.0122															54 Xe 131.29	86 Rn 222
5 Na 22.990	6 Mg 24.305															72 Hf 178.49	84 Po 209
9 F 18.998	10 Ne 20.180															80 Hg 200.59	118 Og 294
11 K 39.098	12 Ca 40.078	13 Sc 44.956	14 Ti 47.88	15 V 50.942	16 Cr 52.00	17 Mn 54.938	18 Fe 55.845	19 Co 58.933	20 Ni 58.69	21 Cu 63.546	22 Zn 65.38	23 Ga 69.723	24 Ge 72.64	25 As 74.922	26 Se 78.96	27 Br 79.904	28 Kr 83.80
19 K 39.098	20 Ca 40.078	21 Sc 44.956	22 Ti 47.88	23 V 50.942	24 Cr 52.00	25 Mn 54.938	26 Fe 55.845	27 Co 58.933	28 Ni 58.69	29 Cu 63.546	30 Zn 65.38	31 Ga 69.723	32 Ge 72.64	33 As 74.922	34 Se 78.96	35 Br 79.904	36 Kr 83.80
37 Rb 85.468	38 Sr 87.62	39 Y 88.906	40 Zr 91.224	41 Nb 92.906	42 Mo 95.94	43 Tc 98.906	44 Ru 101.07	45 Rh 102.91	46 Pd 106.91	47 Ag 107.87	48 Cd 112.41	49 In 114.82	50 Sn 118.71	51 Sb 121.76	52 Te 127.60	53 I 126.91	54 Xe 131.29
55 Cs 132.91	56 Ba 137.33	57 La 138.91	58 Ce 140.12	59 Pr 140.91	60 Nd 144.24	61 Pm 144.91	62 Sm 150.36	63 Eu 151.96	64 Gd 157.25	65 Tb 158.93	66 Dy 162.50	67 Ho 164.93	68 Er 167.26	69 Tm 168.93	70 Yb 173.05	71 Lu 174.97	86 Rn 222
87 Fr 223	88 Ra 226	89 Ac 227	90 Th 232.04	91 Pa 231.04	92 U 238.03	93 Np 237.05	94 Pu 244.06	95 Am 243.06	96 Cm 247.07	97 Bk 247.07	98 Cf 251.08	99 Es 252.08	100 Fm 257.10	101 Md 258.10	102 No 259.10	103 Lr 262.11	118 Og 294
101 Db 262	102 Sg 266	103 Boh 264	104 Hs 277	105 Mt 268	106 Ds 271	107 Rg 272	108 Cn 285	109 Nh 284	110 Fl 289	111 Mc 288	112 Lv 293	113 Ts 294	114 Og 294	115 Uut 289	116 Uuq 291	117 Uub 293	118 Og 294

1	ATOMIC NUMBER
H	SYMBOL
1.008	ATOMIC WEIGHT

- Elemental forms:**
 - Monoatomic:** Noble gases: helium, neon, argon, krypton, xenon, and radon; nonreactive
 - Diatomic:** Gases: fluorine, chlorine, hydrogen, oxygen, and nitrogen; Liquid: bromine; Solid: iodine
 - Molecular solids:** Sulfur
 - Covalent solids:** Carbon (diamond, graphite)
 - Metals:** Most common type of element; high coordination numbers; very dense
- Types of compounds:** Large variety due to ionic, covalent, and metallic bonding options
 - Ionic:** Electron exchange
 - Covalent:** Shared electrons
 - Polar covalent: Unequal sharing
 - Ion complex:** Central ion + Solvent
 - Transition metal compounds:**
 - Coordination compounds: Non-bonding ligands
 - Organometallics: Metal-C bond
 - Complex minerals:** Metal oxides and silicates
- Periodicity:** The chemical properties of an element depend on valence electrons
 - Family of elements:** Column in periodic table
 - The addition of d-orbitals produces an expanded octet

Periodic Trends

- Chemistry determined by the behavior of valence electrons
- Atoms or ions get larger as you go down the column; electron shells are added
- Cation and anions (same charge) decrease in size as the atomic number (Z) increases

- Ionization energies:** Noble gases are the most stable; alkali metals are the easiest to ionize; electron screening lessens Z -effect on orbital energy
- Atomic and ionic radii:** Not real properties; derive from structural data; divide bond into radii using a consistent scheme
 - Ionic radius:** Depends on charge, counter ion, and coordination (crystal type)
 - Atomic radius:** Based on diatomic gas data or metallic solid structural data

Development of the Periodic Table

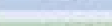
- Missing elements were predicted using periodicity
- Columns 1: Alkali metals
- Columns 2: Alkaline earth metals
- First transition series: Titanium-Copper
- Column 17: Halogens: fluorine, chlorine, bromine, iodine, and astatine
- Column 18: Rare gases: helium, neon, argon, krypton, xenon, and radon
- Artificial: $Z = 90$ or above; radioactive
- Lanthanides: f-subshell elements; starts with $Z = 58$



Practical Application

- The colors of fireworks depend on the elements in the explosive

d-subshell
5 orbitals



Atomic Structure

- Atomic number (Z):** The number of protons in the nucleus; for a neutral atom, Z = number of electrons; for an anion of charge $-q$, $Z = q$ electrons; for a cation of charge $+q$, $Z = q$ electrons
- Mass number (A):** $A = Z + N$, where N = the number of neutrons in the nucleus
- Isotopes:** Atoms with the same Z , different A ; an element can have one or more isotopes; any sample contains a number of isotopes; in practical work, use the average mass number or atomic weight (weighted average of natural isotopes for a given element)

First Ionization Potential vs. Z



Atomic Radius vs. Z



Atomic Quantum Numbers & Orbitals

- n : principle**
 - n : magnetic (orbital direction)
 - n : electron spin
- For each n , the possible l values are $0, 1, \dots, n-1$
- For each l , the possible m_l values are $-l, \dots, 0, \dots, +l$
- m_s has two possible values: $+\frac{1}{2}$ and $-\frac{1}{2}$ (spin up, spin down)
- Each value of n denotes a "shell" in the atomic structure; each l denotes a subshell
 - s-orbital: $l = 0$, 1 type
 - d-orbital: $l = 2$, 5 types
 - p-orbital: $l = 1$, 3 types
 - f-orbital: $l = 3$, 7 types

Many-Electron Atoms

Orbitals and quantum numbers derived for the hydrogen atom are used to describe all many-electron atoms and ions; a more rigorous treatment determines exact energy levels and orbital properties; the main result is that subshells have different energies; the filling of levels is guided by the Aufbau principle:

Aufbau Order for Filling Subshells



Lavc Chemistry Study Guide

L Cohen



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