

## Relative Atomic Mass

### Definition:

The relative atomic mass,  $A_r$ , of an element is defined as the average mass of the isotopes an atom of the element compared with  $\frac{1}{12}$  of the mass of an atom of carbon-12.

Note: \* Symbol for the relative atomic mass is  $A_r$  and  $A_r$  has no unit.

The values of the  $A_r$  of atoms of the elements are given in the Periodic Table and is written on the top of element in the periodic table.

## Atomic Number

- It is the number of protons present in the nucleus of an atom. It is also known as the Proton Number.

## Mass Number/ Nucleon Number

- It is the sum of protons and neutrons present in the nucleus of an atom.

## Relationship between Nucleon number and Relative atomic mass, $A_r$

- The mass of an atom is due mainly to the total number of protons and neutrons in its nucleus (Nucleon number). Thus, we would expect the  $A_r$  of the element to be the same as its nucleon number. *This is true for elements which have no isotopes.*
- For most elements, the nucleon number and the  $A_r$  are not the same due to the *existence of isotopes.*
- The nucleon number is always a whole number (since we can't have fractions of protons or neutrons in the nucleus), but the  $A_r$  values may not be whole numbers. If the  $A_r$  is not a whole number and is different from nucleon number, we can conclude that the element has naturally occurring isotopes.

## Relative Molecular Mass

- Many elements and compounds exist as molecules.
- The mass of a molecule is measured in terms of its relative molecular mass,  $M_r$ .

**Definition:** The relative molecular mass,  $M_r$ , of a substance is defined as the average mass of a molecule of the substance compared with  $\frac{1}{12}$  of the mass of Carbon -12 atom

- Relative molecular mass is often abbreviated to molecular mass.
- Symbol is  $M_r$  and  $M_r$  has no units.
- It is calculated as the sum of the atomic masses of all the atoms in the formula.

# Igcse C1 Paper 2014

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