

## Ideal Gas Law: Multiple Choice

1. The volume of 2.357 mol of oxygen gas at STP is:  
A) 0.1052 L    B) 52.80 L    C) 9.504 L    D) 37.71

ANS: B

1. What is the volume, in liters, occupied by 1.73 moles of  $N_2$  gas at 0.992 atm pressure and a temperature of 75 °C?  
A) 10.7 L    B) 33.8 L    C) 49.8 L    D) 52.2 L

ANS: C

1. A sample of 8 grams of oxygen is at STP ( $1.01 \times 10^5$  Pa and 0 °C). How many molecules are present in the sample?  
A)  $1.51 \times 10^{23}$  molecules    B)  $4.82 \times 10^{23}$  molecules    C)  $3.01 \times 10^{23}$  molecules  
D)  $6.02 \times 10^{23}$  molecules    E) none of the above

ANS: A

1. A certain mass of helium gas occupies 230 mL at 7.5 °C and  $2.3 \times 10^5$  Pa. How many moles are present in the sample?  
A) 0.8484 moles    B) 0.0103 moles    C) 0.0039 moles  
D) 0.0227 moles    E) none of the above

ANS: D

1. What is the number of moles in 500 L of He gas at STP?  
A) 0.05 mol    B) 20.2 mol    C) 22 mol    D) 90 mol

ANS: C

1. What is the number of moles in 500 L of He gas at SATP?  
A) 20.2 mol    B) 0.2 mol    C) 22 mol    D) 40 mol

ANS: A

1. The volume of 1.20 mol of gas at 61.3 kPa and 25.0 °C is \_\_\_\_\_ L.  
A) 135    B) 48.5    C) 52.4    D) 108    E) 55.7

ANS: B

1. Calculate the volume of 12.0 g of helium at 100 °C and 1.2 atm.  
A) 0.013 L    B) 76.5 L    C) 7,750 L    D) 56.0 L

# Ideal Gas Law Multiple Choice Question

**CL Gary**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'CL Gary' text.

## **Ideal Gas Law Multiple Choice Question:**

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