

Key

Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, "PerV-nRT", and the universal gas constant $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}$ to solve the following problems:

If pressure is needed in kPa then convert by multiplying by $101.3 \text{ kPa} / 1 \text{ atm}$ to get
 $R = 8.31 \text{ kPa} \cdot \text{L} / (\text{K} \cdot \text{mole})$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 \text{ atm})(12 \text{ L})}{4 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}}$$

$$T = 204.63 \text{ K}$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of 87°C , how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 \text{ atm} \cdot 31 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 360 \text{ K}}$$

$$n = 1.2586 \text{ mol}$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 400 \text{ K}}{60 \text{ L}}$$

$$P = 1.642 \text{ atm}$$

or $P = 166.29 \text{ kPa}$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56°C , what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 327 \text{ K}}{0.09 \text{ atm}}$$

$$V = 2310.93 \text{ L}$$

- 5) If I have 17 moles of gas at a temperature of 67°C , and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 390 \text{ K}}{88.89 \text{ L}}$$

$$P = 5.34 \text{ atm}$$

or $P = 540.61 \text{ kPa}$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 \text{ atm} \cdot 25 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 300 \text{ K}}$$

$$n = 0.5075 \text{ mol}$$

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Rachel S Tattersall



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constant given by the universal gas constant R and Avogadro's constant N_A . Extra parameters are introduced: a is a measure for the attraction between the particles and b is the average volume excluded from v by a particle. The equation can be cast into the better known form where a is a measure of the attraction between the particles and b is the volume excluded by a mole of particles. A careful distinction.

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1. Understanding the eBook Ideal And Combined Gas Laws Answers Chemfiesta
 - The Rise of Digital Reading Ideal And Combined Gas Laws Answers Chemfiesta
 - Advantages of eBooks Over Traditional Books
2. Identifying Ideal And Combined Gas Laws Answers Chemfiesta
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ideal And Combined Gas Laws Answers Chemfiesta
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ideal And Combined Gas Laws Answers Chemfiesta
 - Personalized Recommendations
 - Ideal And Combined Gas Laws Answers Chemfiesta User Reviews and Ratings
 - Ideal And Combined Gas Laws Answers Chemfiesta and Bestseller Lists

5. Accessing Ideal And Combined Gas Laws Answers Chemfiesta Free and Paid eBooks
 - Ideal And Combined Gas Laws Answers Chemfiesta Public Domain eBooks
 - Ideal And Combined Gas Laws Answers Chemfiesta eBook Subscription Services
 - Ideal And Combined Gas Laws Answers Chemfiesta Budget-Friendly Options
6. Navigating Ideal And Combined Gas Laws Answers Chemfiesta eBook Formats
 - ePub, PDF, MOBI, and More
 - Ideal And Combined Gas Laws Answers Chemfiesta Compatibility with Devices
 - Ideal And Combined Gas Laws Answers Chemfiesta Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ideal And Combined Gas Laws Answers Chemfiesta
 - Highlighting and Note-Taking Ideal And Combined Gas Laws Answers Chemfiesta
 - Interactive Elements Ideal And Combined Gas Laws Answers Chemfiesta
8. Staying Engaged with Ideal And Combined Gas Laws Answers Chemfiesta
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ideal And Combined Gas Laws Answers Chemfiesta
9. Balancing eBooks and Physical Books Ideal And Combined Gas Laws Answers Chemfiesta
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ideal And Combined Gas Laws Answers Chemfiesta
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ideal And Combined Gas Laws Answers Chemfiesta
 - Setting Reading Goals Ideal And Combined Gas Laws Answers Chemfiesta
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ideal And Combined Gas Laws Answers Chemfiesta
 - Fact-Checking eBook Content of Ideal And Combined Gas Laws Answers Chemfiesta
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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