

Question			Answer	Marks	Guidance
2	(a)	(i)	$Al^{3+}$ ✓ $SO_4^{2-}$ ✓	2	
		(ii)	$Al_2O_3(s) + 3H_2SO_4(aq) \rightarrow Al_2(SO_4)_3(aq) + 3H_2O(l)$ Correct species <b>AND</b> correctly balanced ✓ state symbols on correct species ✓	2	<b>ALLOW</b> multiples
		(iii)	(The number of) water(s) of crystallisation ✓	1	<b>IGNORE</b> hydrated <b>OR</b> hydrous <b>OR</b> 'contains water'
		(iv)	First check the answer on the answer line. If answer = 16, award 3 marks  Correctly calculates amount of $Al_2(SO_4)_3$ : $6.846 / 342.3 = 0.02(00)$ mol ✓  Correctly calculates amount of $H_2O$ : $5.760 / 18.0 = 0.32(0)$ mol ✓  Correctly calculates whole number ratio of mol of $H_2O$ : $Al_2(SO_4)_3$ to give $x = 16$ ✓	3	If there is an alternative answer, check to see if there is any ECF credit possible using working below  <b>ALLOW</b> as ECF from $12.606/342.3 = 0.0368(273)$ <b>AND</b> $0.32/0.0368(273)$ To give $x = 9$ for two marks  <b>ALLOW</b> calculator value or rounding to 2 significant figures or more <b>BUT IGNORE</b> 'trailing' zeroes, eg 0.200 allowed as 0.2. <b>ALLOW</b> ECF for calculation of correctly rounded <b>whole</b> number value of $H_2O$ from incorrect mol of $H_2O$ and / or incorrect mol of $Al_2(SO_4)_3$ <b>BUT</b> $x$ must be a whole number  <b>ALLOW</b> alternative method Mol of $Al_2(SO_4)_3$ : $6.846 / 342.3 = 0.02(00)$ mol (first mark)  Molar mass of $Al_2(SO_4)_3 \cdot xH_2O$ : $12.606 / 0.02(00) = 630.3$ g mol <sup>-1</sup> (second mark)  Mass of water per mol = $630.3 - 342.3 = 288$ <b>AND</b> $288/18$ to give $x = 16$ (third mark)

# Mark Scheme June 2013 Core One Ocr

**Colin Shephard, Rosemary Rees**



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