## Unit 6a Review questions HL (answers at end)

#### May 2010 HL p2 TZ1

Brass is a copper containing alloy with many uses. An analysis is comied out to determine the percentage of copper present in three identical samples of brass. The reactions involved in this analysis are shown below.

- Step 1:  $Cu(s) * 2HDO_s(aq) * 2H^*(aq) \rightarrow Cu^{2*}(aq) * 2NO_s(q) * 2H_sO(b)$
- Step 2:  $4\Gamma(\log) + 2Cu^{2r}(\log) \rightarrow 2Cul(s) 1, (\log)$
- Step J:  $I_{*}(mq) = 2S_{*}O_{*}^{-1}(mq) \rightarrow 2I_{*}(mq) + S_{*}O_{*}^{-1}(mq)$
- (ii) Deduce the change in the oxidation numbers of copper and nitrogen in step 1. /2/

Coppert:

Nitragera:

(iii) Identify the existicing agent in way 1. (11)

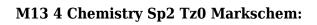
(b) A student carried out this experiment three times, with three identical small beans sails, and obtained the following results.

Mass of brass = 0.456 g. 2:40.001 g.

Titre	1	2	3
Initial volume of 0.100 and $dm^2 S_j O_j^{2-} (\pm 0.05 cm^2)$	0.00	0.00	0.00
Final volume of 0.100 mol dm" S <sub>2</sub> O <sub>2</sub> " (±0.05 cm")	28.50	28.60	28.40
Volume added of 0.100 mol dm "S <sub>j</sub> O <sub>j</sub> " (±0.10 cm <sup>2</sup> )	28.50	28,66	28.40
Average volume added of 0.100 mol $dm^{-1} S_2O_4^{-1}$ ( $\pm 0.10 cm^2$ )	28.50		

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**DP Hallahan** 



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