EXAMINATION DATA SHEET FOR THE PHYSICAL SCIENCES (CHEMISTRY)

TABLE 1 PHYSICAL CONSTANTS

NAME	SYMBOL	VALUE
Magnitude of charge on electron	е	1,6 × 10 ⁻¹⁹ C
Mass of an electron	me	9,1 × 10 ⁻³¹ kg
Standard pressure	p⁰	1,01 × 10 ⁵ Pa
Motar gas volume at STP	V_	22,4 dm ³ · mol-1
Standard temperature	T*	273 K
Avogadro's constant	NA	6,02 × 10 ²⁹ mol ⁻¹
Faraday's constant	F	96 500 C- mol-1

TABLE 2 CHEMISTRY FORMULAE

$$n = \frac{m}{M}$$

$$n = \frac{N}{N_A}$$

$$r = \frac{V}{V_m}$$

$$c = \frac{n}{V} \text{ OR } c = \frac{m}{MV}$$

$$K_w = [H_3O^+] \text{ [OHT]} = 1 \times 10^{-14}$$

$$\text{at 25 °C (298 K)}$$

$$q = nF$$

$$E_{cell}^\theta = E_{cathode}^\theta - E_{anode}^\theta$$

$$E_{cell}^\theta = E_{oxidising agent}^\theta - E_{reducing agent}^\theta$$

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