

Experiment 1: Charles's Law

Table 1: Temperature and Volume Data

Temperature Conditions	Temperature ($^{\circ}\text{C}$)	Temperature (K)	Syringe Volume (mL)
Room Temperature	23 C	296.15 K	5 ml
Hot Water	45 C	318.15 K	1 ml
Ice Water	5 C	278.15 K	0.3 ml

Questions:

1. Create a graph of temperature and volume data. Place temperature on the x-axis (in Kelvin) and volume (mL) on the y-axis. Insert your graph below.
2. What happened to the volume of gas when the syringe was exposed to various temperature conditions? Using the concepts explored in the Introduction, describe why this occurred, keeping in mind the definition of temperature.
3. Draw a straight line of best fit through your data points on your graph in question 1, extrapolating the line until it intersects the (negative) x-axis. Why can you assume a linear relationship (a straight-lined slope)?
4. At what temperature does your line intersect the x-axis? What volume corresponds to this temperature?

Ideal Gas Law Lab Escience Answers

Otto Maass



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