

Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The charge on an ion is
a. always positive.
b. always negative.
c. either positive or negative.
d. zero.
- _____ 2. According to the octet rule, a calcium atom has a tendency to
a. lose one electron.
b. lose two electrons.
c. gain one electron.
d. gain two electrons.
- _____ 3. If a compound forms by ionic bonding, which is *not* true?
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
b. The net charge of the compound is zero.
c. The compound contains just two atoms, each of opposite charge.
d. Several ions group together in a tightly packed structure.
- _____ 4. The only property listed that is *not* characteristic of ionic compounds is
a. high melting point.
b. hardness.
c. lack of crystal structure.
d. good conductor of electricity.
- _____ 5. Which formula listed below represents a polyatomic ion?
a. NH_4^+
b. H_2SO_4
c. NaCl
d. H_2O
- _____ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
a. ionic substances tend to vaporize at room temperature.
b. ionic substances are brittle.
c. attractive forces between ions are greater than the attractive forces between molecules.
d. the numbers of positive and negative charges are equal in an ionic compound.

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Table of Contents Holt Modern Chemistry Chapter 12 Review Answers

1. Understanding the eBook Holt Modern Chemistry Chapter 12 Review Answers
 - The Rise of Digital Reading Holt Modern Chemistry Chapter 12 Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Modern Chemistry Chapter 12 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Modern Chemistry Chapter 12 Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Modern Chemistry Chapter 12 Review Answers
 - Personalized Recommendations
 - Holt Modern Chemistry Chapter 12 Review Answers User Reviews and Ratings
 - Holt Modern Chemistry Chapter 12 Review Answers and Bestseller Lists
5. Accessing Holt Modern Chemistry Chapter 12 Review Answers Free and Paid eBooks
 - Holt Modern Chemistry Chapter 12 Review Answers Public Domain eBooks
 - Holt Modern Chemistry Chapter 12 Review Answers eBook Subscription Services
 - Holt Modern Chemistry Chapter 12 Review Answers Budget-Friendly Options

6. Navigating Holt Modern Chemistry Chapter 12 Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Holt Modern Chemistry Chapter 12 Review Answers Compatibility with Devices
 - Holt Modern Chemistry Chapter 12 Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Modern Chemistry Chapter 12 Review Answers
 - Highlighting and Note-Taking Holt Modern Chemistry Chapter 12 Review Answers
 - Interactive Elements Holt Modern Chemistry Chapter 12 Review Answers
8. Staying Engaged with Holt Modern Chemistry Chapter 12 Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Modern Chemistry Chapter 12 Review Answers
9. Balancing eBooks and Physical Books Holt Modern Chemistry Chapter 12 Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Modern Chemistry Chapter 12 Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Holt Modern Chemistry Chapter 12 Review Answers
 - Setting Reading Goals Holt Modern Chemistry Chapter 12 Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Holt Modern Chemistry Chapter 12 Review Answers
 - Fact-Checking eBook Content of Holt Modern Chemistry Chapter 12 Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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