

Name: \_\_\_\_\_

**CHEMISTRY 101**  
**EXAM 3 FORM B****SECTIONS 501-511****SPRING 2005**  
**DR. KEENEY-KENNICUTT**

- Directions:
- (1) Put your name and signature on PART 2 of the exam where indicated.
  - (2) Sign the Aggie Code on PART 2 of this exam.
  - (3) Each multiple choice question is actually 2 questions on your scanning sheet. If you are sure of an answer, put the same answer down for both questions for 5 pts. If you cannot decide between two answers, put one answer down for one question and the other answer down for the other question. If you get one correct you'll get half credit for 2.5 pts. If there is an ambiguous multiple choice question, use the last page to explain your answer.
  - (4) Do NOT write on the envelope.
  - (5) When finished, put everything in the envelope and wait to be excused. At the table, take everything out of the envelope. You can pick up the multiple choice part with the answers outside my office after the exam.
  - (6) There are a total of 27 questions (15 actual questions).

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**PART 1**

---

**1&2.** The number of protons and electrons found in the aluminum ion that is isoelectronic with a noble gas is:

- (a) 13 p and 16 e<sup>-</sup>                      (b) 10 p and 13 e<sup>-</sup>                      (c) 13 p and 10 e<sup>-</sup>  
(d) 13 p and 13 e<sup>-</sup>                      (e) some other combination

**3&4.** Which of the following bonds can be classified as polar covalent?

- (a) Cu-Ti                      (b) H-F                      (c) S-I                      (d) C-S                      (e) Na-F

**5&6.** Which indication of relative acid strength is WRONG?

- (a)  $\text{H}_2\text{S} < \text{H}_2\text{Se}$                       (b)  $\text{H}_2\text{SeO}_4 < \text{H}_2\text{SO}_4$                       (c)  $\text{HBrO}_2 > \text{HBrO}_3$   
(d)  $\text{H}_2\text{SO}_4 > \text{HSO}_4^-$                       (e)  $\text{HF} < \text{HI}$

# Guide For Chemistry Spring Exam

**Liying Dong**



## **Guide For Chemistry Spring Exam:**

This book delves into Guide For Chemistry Spring Exam. Guide For Chemistry Spring Exam is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Guide For Chemistry Spring Exam, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Guide For Chemistry Spring Exam
- Chapter 2: Essential Elements of Guide For Chemistry Spring Exam
- Chapter 3: Guide For Chemistry Spring Exam in Everyday Life
- Chapter 4: Guide For Chemistry Spring Exam in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Guide For Chemistry Spring Exam. This chapter will explore what Guide For Chemistry Spring Exam is, why Guide For Chemistry Spring Exam is vital, and how to effectively learn about Guide For Chemistry Spring Exam.

3. In chapter 2, the author will delve into the foundational concepts of Guide For Chemistry Spring Exam. The second chapter will elucidate the essential principles that need to be understood to grasp Guide For Chemistry Spring Exam in its entirety.

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6. In chapter 5, the author will draw a conclusion about Guide For Chemistry Spring Exam. This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Guide For Chemistry Spring Exam.

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