



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the reaction $\rightleftharpoons$.

What is Equilibrium? Equilibrium is a state of balance between the forward and reverse reactions. At equilibrium, the concentrations of the chemical species remain the same with time. At equilibrium, the rate of the forward reaction is equal to the rate of the reverse reaction. Equilibrium is a dynamic process. It is not a static process. Equilibrium Constant, K_c The equilibrium constant, K_c, is a measure of the extent of a chemical reaction. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant, K_c, is defined as: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ where $[A]$, $[B]$, $[C]$, and $[D]$ are the concentrations of the chemical species at equilibrium.	Equilibrium Constant For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant, K_c, is defined as: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ where $[A]$, $[B]$, $[C]$, and $[D]$ are the concentrations of the chemical species at equilibrium.	Concentration Time Reactants: $[A]_0, [B]_0 \rightarrow [A]_{eq}, [B]_{eq}$ Products: $[C]_0, [D]_0 \rightarrow [C]_{eq}, [D]_{eq}$																
Equilibrium Constant, K_c The equilibrium constant, K_c, is a measure of the extent of a chemical reaction. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant, K_c, is defined as: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ where $[A]$, $[B]$, $[C]$, and $[D]$ are the concentrations of the chemical species at equilibrium. Equilibrium Constant, K_p The equilibrium constant, K_p, is a measure of the extent of a chemical reaction. It is defined as the ratio of the partial pressures of the products to the partial pressures of the reactants, each raised to the power of its stoichiometric coefficient. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant, K_p, is defined as: $K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$ where P_A, P_B, P_C, and P_D are the partial pressures of the chemical species at equilibrium.	<table><tr><th>Equilibrium Constant, K_c</th><th>Meaning</th></tr><tr><td>$K_c > 1$</td><td>Concentration of products is greater than concentration of reactants.</td></tr><tr><td>$K_c = 1$</td><td>Concentration of products is equal to concentration of reactants.</td></tr><tr><td>$K_c < 1$</td><td>Concentration of reactants is greater than concentration of products.</td></tr></table>	Equilibrium Constant, K_c	Meaning	$K_c > 1$	Concentration of products is greater than concentration of reactants.	$K_c = 1$	Concentration of products is equal to concentration of reactants.	$K_c < 1$	Concentration of reactants is greater than concentration of products.									
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