

The isotope fractionation factor

The isotope fractionation factor is defined as: $\alpha_b^a = \frac{R_a}{R_b}$

where R_A , R_B are the isotope ratios in two phases (e.g., carbonate and water, or water vapor and water)

For example, consider: $\text{H}_2\text{O}(l) \leftrightarrow \text{H}_2\text{O}(v)$ at 25°C

$$\alpha_v^l(O) = \frac{R_l}{R_v} = \frac{\left(^{18}\text{O}/^{16}\text{O}\right)_l}{\left(^{18}\text{O}/^{16}\text{O}\right)_v} = 1.0092$$

The differences in isotope ratios are relatively small and are expressed as parts per thousand (per mil) deviations from a standard.

Fractionation Formula Examples

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