

## Übung 1: Semester 1 - Vorgehen - 1. und 2. Semester

### 1. Semester - Vorgehen

Die Aufgabenstellung ist in der Aufgabenstellung des 1. Semesters zu finden. Die Aufgabenstellung ist in der Aufgabenstellung des 1. Semesters zu finden.

### 2. Semester - Vorgehen

Die Aufgabenstellung ist in der Aufgabenstellung des 2. Semesters zu finden. Die Aufgabenstellung ist in der Aufgabenstellung des 2. Semesters zu finden.

### 3. Semester - Vorgehen

Die Aufgabenstellung ist in der Aufgabenstellung des 3. Semesters zu finden. Die Aufgabenstellung ist in der Aufgabenstellung des 3. Semesters zu finden.

### 4. Semester - Vorgehen

Die Aufgabenstellung ist in der Aufgabenstellung des 4. Semesters zu finden. Die Aufgabenstellung ist in der Aufgabenstellung des 4. Semesters zu finden.

### 5. Semester - Vorgehen

Die Aufgabenstellung ist in der Aufgabenstellung des 5. Semesters zu finden. Die Aufgabenstellung ist in der Aufgabenstellung des 5. Semesters zu finden.

### 6. Semester - Vorgehen

Die Aufgabenstellung ist in der Aufgabenstellung des 6. Semesters zu finden. Die Aufgabenstellung ist in der Aufgabenstellung des 6. Semesters zu finden.

### 7. Semester - Vorgehen

Die Aufgabenstellung ist in der Aufgabenstellung des 7. Semesters zu finden. Die Aufgabenstellung ist in der Aufgabenstellung des 7. Semesters zu finden.

### 8. Semester - Vorgehen

### 9. Semester - Vorgehen

Die Aufgabenstellung ist in der Aufgabenstellung des 9. Semesters zu finden. Die Aufgabenstellung ist in der Aufgabenstellung des 9. Semesters zu finden.

### 10. Semester - Vorgehen

Die Aufgabenstellung ist in der Aufgabenstellung des 10. Semesters zu finden. Die Aufgabenstellung ist in der Aufgabenstellung des 10. Semesters zu finden.

Die Aufgabenstellung ist in der Aufgabenstellung des 11. Semesters zu finden. Die Aufgabenstellung ist in der Aufgabenstellung des 11. Semesters zu finden.

### 12. Semester - Vorgehen

Die Aufgabenstellung ist in der Aufgabenstellung des 12. Semesters zu finden. Die Aufgabenstellung ist in der Aufgabenstellung des 12. Semesters zu finden.

### 13. Semester - Vorgehen

Die Aufgabenstellung ist in der Aufgabenstellung des 13. Semesters zu finden. Die Aufgabenstellung ist in der Aufgabenstellung des 13. Semesters zu finden.

### 14. Semester - Vorgehen

Die Aufgabenstellung ist in der Aufgabenstellung des 14. Semesters zu finden. Die Aufgabenstellung ist in der Aufgabenstellung des 14. Semesters zu finden.

### 15. Semester - Vorgehen

# Everfi Module 9 Final Quiz

**Christian G. Meyer**

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web answer  $\text{Na}^+ \text{aq}$  and  $\text{SO}_4^{2-} \text{aq}$  identify the chemical equation the spectator ions  $\text{KOH} \text{aq}$   $\text{HNO}_3$   $\text{KNO}_3 \text{aq}$   $\text{H}_2\text{O}$   $\text{l}$   $\text{K}^+ \text{aq}$   $\text{OH}^- \text{aq}$   $\text{H}^+ \text{aq}$

no3 aq k aq no3 aq h2o l answer k aq and no3 aq identify the chemical equation the spectator ions pb no3 2 aq 2ki aq pbi2 s 2kno3 aq

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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction 2ag aq cr 2o 7 2 aq rightarrow ag 2cr 2o 7 s label 4 2 3

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water  $\text{H}^+ \text{OH}^- \rightarrow \text{H}_2\text{O}$  the spectator ions in  $\text{KOH} \text{aq}$  and  $\text{HNO}_3 \text{aq}$  are  $\text{K}^+ \text{aq}$  and  $\text{NO}_3^- \text{aq}$  how to identify and list the

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web the  $\text{Ag}^+ \text{aq}$  and  $\text{Cl}^- \text{aq}$  ions become  $\text{AgCl s}$  but the  $\text{Na}^+ \text{aq}$  ions and the  $\text{NO}_3^- \text{aq}$  ions stay as  $\text{Na}^+ \text{aq}$  ions and  $\text{NO}_3^- \text{aq}$  ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be  $2\text{Cl}^- + \text{Cu}^{2+} \rightarrow \text{CuCl}_2$  these spectator ions are not actively involved in the reaction the existence of spectator ions affects the debye length