

## Electron Configurations - KEY

*Note: The electron configurations in this worksheet assume that lanthanum (La) is the first element in the 4f block and that actinium (Ac) is the first element in the 5f block. If your periodic table doesn't agree with this, your answers for elements near the f-orbitals may be slightly different (one more/less electron for the f-orbital).*

- 1) sodium  $1s^2 2s^2 2p^6 3s^1$
- 2) iron  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^5$   
also might be written (due to the close energy of d and f and similarities of elements versus predictable)  
memorized config:  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 6d^1 5f^4$
- 6) cobalt  $[Ar] 4s^2 3d^7$
- 7) silver  $[Kr] 5s^2 4d^9$
- 8) tellurium  $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium  $[Rn] 7s^2$
- 10) lawrencium  $[Rn] 7s^2 5f^{14} 6d^1$   
also might be written (due to the close energy of d and f and similarities of elements versus predictable)  
memorized config:  $[Rn] 7s^2 6d^1 5f^{14}$
- 11)  $1s^2 2s^2 2p^6 3s^2 3p^4$  sulfur
- 12)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$  rubidium
- 13)  $[Kr] 5s^2 4d^{10} 5p^3$  antimony
- 14)  $[Xe] 6s^2 4f^{14} 5d^6$  osmium
- 15)  $[Rn] 7s^2 5f^{11}$  einsteinium note: no d's, so the p. table referenced has Ac as 5f
- 16)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$  not valid (take a look at "4d"; must be 3 d not 4!)
- 17)  $1s^2 2s^2 2p^6 3s^3 3d^5$  not valid (3p comes after 3s); also no 3 electrons in s
- 18)  $[Ra] 7s^2 5f^8$  not valid (radium isn't a noble gas)
- 19)  $[Kr] 5s^2 4d^{10} 5p^5$  valid
- 20)  $[Xe]$  not valid (an element can't be its own electron configuration)

# Electron Configuration Section 5 Answer Key

**Atef Korchef**



## **Electron Configuration Section 5 Answer Key:**

Understanding General Chemistry Atef Korchef, 2022-03-07 Understanding General Chemistry details the fundamentals of general chemistry through a wide range of topics relating the structure of atoms and molecules to the properties of matter. Written in an easy to understand format with helpful pedagogy to fuel learning the book features main objectives at the beginning of each chapter, get smart sections and check your reading section at the end of each chapter. The text is filled with examples and practices that illustrate the concepts at hand. In addition, a summary and extensive MCQs exercises and problems with the corresponding answers and explanations are readily available. Additional features include Alerts students to common mistakes and explains in simple ways and clear applications how to avoid these mistakes. Offers answers and comments alongside sample problems enabling students to self evaluate their skill level. Includes powerful methods easy steps simple and accurate interpretations and engaging applications to help students understand complex principles. Provides a bridge to more complex topics such as solid state chemistry, organometallic chemistry, chemistry of main group elements, inorganic chemistry and physical chemistry. This introductory textbook is ideal for chemistry courses for non science majors as well as health sciences and preparatory engineering students.

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