

Name _____ Class _____ Date _____

Assessment

Chapter Test B

Atomic Physics

MULTIPLE CHOICE

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. What is the frequency of a photon with an energy of $1.00 \times 10^{-19} \text{ J}$?
($h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s}$)
a. $1.00 \times 10^{14} \text{ Hz}$ c. $3.00 \times 10^{12} \text{ Hz}$
b. $2.00 \times 10^{14} \text{ Hz}$ d. $4.00 \times 10^{12} \text{ Hz}$
- _____ 2. Light with an energy equal to three times the work function of a given metal causes the metal to eject photoelectrons. What is the ratio of the maximum photoelectron kinetic energy to the work function?
a. 1 : 1 c. 3 : 1
b. 2 : 1 d. 4 : 1
- _____ 3. A monochromatic light beam with a quantum energy value of 3.0 eV is incident upon a photocell. The work function of the photocell is 1.6 eV. What is the maximum kinetic energy of the ejected electrons?
a. 4.6 eV c. 1.4 eV
b. 4.8 eV d. 2.4 eV
- _____ 4. What causes the bright lines in the emission spectrum of an element to occur?
a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
d. Photons are emitted when electrons jump from a lower-energy to a higher-energy state.
- _____ 5. What causes the dark lines in the absorption spectrum of an element to occur?
a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
d. Photons are emitted when electrons jump from a lower-energy to a higher-energy state.

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Clifford Lane



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1. Understanding the eBook Holt Physical Science Ch 14 Test
 - The Rise of Digital Reading Holt Physical Science Ch 14 Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Physical Science Ch 14 Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Physical Science Ch 14 Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Physical Science Ch 14 Test
 - Personalized Recommendations
 - Holt Physical Science Ch 14 Test User Reviews and Ratings
 - Holt Physical Science Ch 14 Test and Bestseller Lists
5. Accessing Holt Physical Science Ch 14 Test Free and Paid eBooks
 - Holt Physical Science Ch 14 Test Public Domain eBooks
 - Holt Physical Science Ch 14 Test eBook Subscription Services
 - Holt Physical Science Ch 14 Test Budget-Friendly Options
6. Navigating Holt Physical Science Ch 14 Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Holt Physical Science Ch 14 Test Compatibility with Devices
 - Holt Physical Science Ch 14 Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Physical Science Ch 14 Test
 - Highlighting and Note-Taking Holt Physical Science Ch 14 Test
 - Interactive Elements Holt Physical Science Ch 14 Test
8. Staying Engaged with Holt Physical Science Ch 14 Test

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Physical Science Ch 14 Test
9. Balancing eBooks and Physical Books Holt Physical Science Ch 14 Test
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Physical Science Ch 14 Test
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Holt Physical Science Ch 14 Test
- Setting Reading Goals Holt Physical Science Ch 14 Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Holt Physical Science Ch 14 Test
- Fact-Checking eBook Content of Holt Physical Science Ch 14 Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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