

Chemical Bonding

Key Concepts

- Section 1: Introduction to Chemical Bonding**
 What is a chemical bond and why does it form?
- Section 2: Chemical Bonding and Molecular Compounds**
 What is a molecular formula?
 What are the characteristics of a covalent bond?
 How does a Lewis structure represent a molecule?
- Section 3: Ionic Bonding and Ionic Compounds**
 What is a crystal lattice?
 How are ionic and molecular compounds different?
- Section 4: Metallic Bonding**
 How does metallic bonding affect the properties of metals?
- Section 5: Molecular Geometry**
 What is VSEPR theory and how is it used?
 What other factors affect the shape of covalent molecules?

LEARNING OBJECTIVES
 LO 1.1.1: Define and explain how the Lewis dot model represents chemical bonding.
 LO 1.1.2: Draw a Lewis dot model showing valence electrons for a molecule or ionic compound, and explain how it represents the compound.
 LO 1.1.3: Identify the number of bonds an atom can form and the number of valence electrons.
 LO 1.1.4: Explain the difference in properties between ionic and molecular compounds.
 LO 1.1.5: Explain the difference in properties between ionic and molecular compounds.
 LO 1.1.6: Explain the difference in properties between ionic and molecular compounds.
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 LO 1.1.9: Explain the difference in properties between ionic and molecular compounds.
 LO 1.1.10: Explain the difference in properties between ionic and molecular compounds.

Review Previous Concepts

1. What are valence electrons?

2. What quantity does the electronegativity of an atom represent?

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Léo Pomar



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