

10th Grade Honors Chemistry Final Exam Flashcards With Complete Solutions

Thermochemistry CORRECT ANSWERS The study of heat effects in chemical reactions.

Exothermic Reaction CORRECT ANSWERS A reaction that releases energy in the form of heat.

Examples: Combustion, freezing water.

Heat generated is added to the product side (right side) because heat is released.

Endothermic Reaction CORRECT ANSWERS A reaction that absorbs energy in the form of heat.

Examples: Melting ice, evaporating liquid water, boiling water.

Heat generated is added to the reactant side (left side) because heat is added.

Enthalpy (H or ΔH) CORRECT ANSWERS The energy after taking into account the pressure and volume of a substance.

If pressure remains constant, then heat absorbed or released during a chemical reaction is equal to the enthalpy change for the reaction.

$\Delta H =$ CORRECT ANSWERS $H_{\text{Products}} - H_{\text{Reactants}}$.

Standard State of Element (Temperature and Pressure) CORRECT ANSWERS
Temperature = 25° Celsius.
Pressure = 1 atm.

ΔH° CORRECT ANSWERS Standard enthalpy change = the heat transferred in a reaction with reactants and products in standard states.

Determining Enthalpy Change CORRECT ANSWERS $\Delta H = (-)$ Exo.
 $\Delta H = (+)$ Endo.

If heat is released as in an exothermic process the temperature will increase, but the product will have less energy than the reactant so enthalpy will be negative.

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Mary Millhollon, Richard H. Langley



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