

Chapter 2 Properties of Matter

Section 2.1 Classifying Matter

(pages 38–44)

This section explains how materials are classified as pure substances or mixtures. It discusses types of pure substances and mixtures.

Reading Strategy (page 38)

Summarizing As you read, complete the classification of matter in the diagram below. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

**Pure Substances (page 39)**

- Is the following sentence true or false? Every sample of a pure substance has exactly the same composition and the same properties. true
- What are the two categories of pure substances?
a. Elements b. Compounds

Elements (pages 39–40)

- What is an element? An element is a substance that cannot be broken down into simpler substances.
- Is the following sentence true or false? The smallest particle of an element is an atom. true
- Why does an element have a fixed, uniform composition? An element has a fixed composition because it contains only one type of atom.
- Circle the letter before each element that is a gas at room temperature.
a. carbon b. oxygen
c. mercury d. nitrogen

Match each element to its correct symbol.

Element	Symbol
<u>b.</u> 7. aluminum	a. C
<u>c.</u> 8. gold	b. Al
<u>a.</u> 9. carbon	c. Au

Matter Physical Science Study Guide Answers

Brian W. Stewart



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Chemical News and Journal of Physical Science William Crookes, James H. Gardiner, Gerald Druce, H. W. Blood-Ryan, 1916

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