Limiting Reagent & Percent Yield Practice Worksheet

- When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.
 - a. Write the balanced equation for the reaction given above:

 $CuCl_2 + NaNO_3 \rightarrow Cu(NO_3)_2 + NaCl$ $CuCl_2 + 2 NaNO_3 \rightarrow Cu(NO_3)_2 + 2 NaCl$

b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

15 g CuCl₂ x 1 mol CuCl₂ x 2 mol NaCl x 58.443 g NaCl = 13 g NaCl 134.362g CuCl₂ 1 mol CuCl₂ 1 mol NaCl

20 g NaNO₃ x <u>1 mol NaNO₃</u> x <u>2 mol NaCl</u> x <u>58.443 g NaCl</u> = 13.8 g NaCl 84.997g NaNO₃ 2 mol NaNO₃ 1 mol NaCl

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? copper (11) chloride
- d. How many grams of copper (II) nitrate is formed?

15g CuCl₂ x 1 mol CuCl₂ x 1 mol Cu(NO₃)₂ x 187.467 g Cu(NO₃)₂ = 20.9 g Cu(NO₃)₂ = 134.452g CuCl₂ 1 mol CuCl₂ 1 mol Cu(NO₃)₂

e. How much of the excess reagent is left over in this reaction?

15g CuCl₂ x 1 mol CuCl₂ x 2 mol NaNO₃ x 84,997g NaNO₃ = 19 g NaNO₃ 134,452g CuCl₂ 1 mol CuCl₂ 1 mol NaNO₃

20 g NaNO₅ - 19 g NaNO₅ = 1 g NaNO₅

f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction? Percent yield = actual yield / theoretical yield × 100 percent.

11.3/13.0 x 100% = 86.9%

- When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.
 - a. Balance the following equation:

 $Pb(NO_3)_{2 \text{ (sq)}} + NaI_{\text{ (sq)}} \rightarrow PbI_{2 \text{ (s)}} + NaNO_{3 \text{ (sq)}}$ $Pb(NO_3)_{2 \text{ (sq)}} + 2 NaI_{\text{ (sq)}} \rightarrow PbI_{2 \text{ (s)}} + 2 NaNO_{3 \text{ (sq)}}$

b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

25.0 g Pb(NO₃)₂ x 1 mol Pb(NO₃)₂ x 2 mol NaNO₃ x 84.997g NaNO₃ = 12.8 g NaNO₃ 331.214 g Pb(NO₃)₂ 1 mol Pb(NO₃)₂ 1 mol NaNO₃

15.0 g Nal x 1 mol Nal x 2 mol NaNO₃ x 84.997g NaNO₃ = 8.51 g NaNO₃ 149.89g Nal 2 mol Nal 1 mol NaNO₃

8.51 g NaNO3 can be made

<u>Limiting Reagent And Percent Yield Worksheet Ib</u> <u>Chemistry</u>

Victor M. Corman

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Table of Contents Limiting Reagent And Percent Yield Worksheet Ib Chemistry

- 1. Understanding the eBook Limiting Reagent And Percent Yield Worksheet Ib Chemistry
 - o The Rise of Digital Reading Limiting Reagent And Percent Yield Worksheet Ib Chemistry
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Limiting Reagent And Percent Yield Worksheet Ib Chemistry
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reagent And Percent Yield Worksheet Ib Chemistry
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Limiting Reagent And Percent Yield Worksheet Ib Chemistry
 - Personalized Recommendations
 - Limiting Reagent And Percent Yield Worksheet Ib Chemistry User Reviews and Ratings

- Limiting Reagent And Percent Yield Worksheet Ib Chemistry and Bestseller Lists
- 5. Accessing Limiting Reagent And Percent Yield Worksheet Ib Chemistry Free and Paid eBooks
 - Limiting Reagent And Percent Yield Worksheet Ib Chemistry Public Domain eBooks
 - Limiting Reagent And Percent Yield Worksheet Ib Chemistry eBook Subscription Services
 - Limiting Reagent And Percent Yield Worksheet Ib Chemistry Budget-Friendly Options
- 6. Navigating Limiting Reagent And Percent Yield Worksheet Ib Chemistry eBook Formats
 - o ePub, PDF, MOBI, and More
 - Limiting Reagent And Percent Yield Worksheet Ib Chemistry Compatibility with Devices
 - Limiting Reagent And Percent Yield Worksheet Ib Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reagent And Percent Yield Worksheet Ib Chemistry
 - Highlighting and Note-Taking Limiting Reagent And Percent Yield Worksheet Ib Chemistry
 - Interactive Elements Limiting Reagent And Percent Yield Worksheet Ib Chemistry
- 8. Staying Engaged with Limiting Reagent And Percent Yield Worksheet Ib Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reagent And Percent Yield Worksheet Ib Chemistry
- 9. Balancing eBooks and Physical Books Limiting Reagent And Percent Yield Worksheet Ib Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reagent And Percent Yield Worksheet Ib Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Limiting Reagent And Percent Yield Worksheet Ib Chemistry
 - Setting Reading Goals Limiting Reagent And Percent Yield Worksheet Ib Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Limiting Reagent And Percent Yield Worksheet Ib Chemistry
 - Fact-Checking eBook Content of Limiting Reagent And Percent Yield Worksheet Ib Chemistry
 - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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