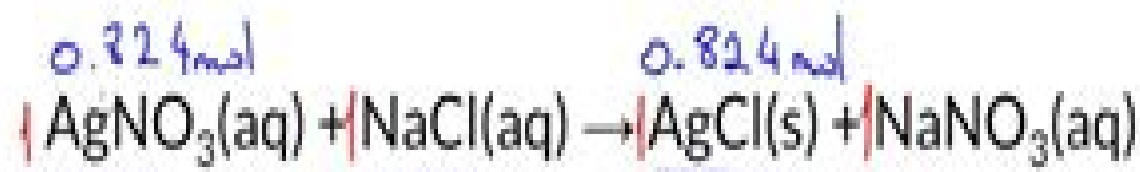


When 140.0 g of AgNO_3 was added to an aqueous solution of NaCl , 86.0 g of AgCl was collected as a white precipitate. Which salt was the limiting reactant in this reaction? How many grams of NaCl was present in the solution?



1) L.R. - ?

2) $m(\text{NaCl})$ - ?

$$n(\text{AgNO}_3) = \frac{m}{M} = \frac{140 \text{ g}}{169.9 \text{ g/mol}} = 0.824 \text{ mol}$$

$$n(\text{AgCl}) = \frac{m}{M} = \frac{86.0 \text{ g}}{143.3 \text{ g/mol}} = \underline{0.600 \text{ mol A.Y.}}$$

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General Organic and Biological Chemistry Kenneth W. Raymond, 2009-12-14 This general organic and biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry Students need have no previous background in chemistry but should possess basic math skills The text features numerous helpful problems and learning features

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