

Boyle's Law: A sample of N_2 gas at 101 kPa occupies 1.7 L. What volume will the same sample occupy if the pressure is decreased to 31 kPa at constant temperature?

$$P_1 V_1 = P_2 V_2 \Rightarrow V_2 = V_1 \times \frac{P_1}{P_2} = 1.7 \text{ L} \times \frac{101 \text{ kPa}}{31 \text{ kPa}} = 5.5 \text{ L}$$

Charles' Law: A sample of CO_2 gas occupies 551 mL at a temperature of 25.0°C. What volume will the sample occupy if the temperature is increased to 100.0°C at constant pressure?

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \Rightarrow V_2 = V_1 \times \frac{T_2}{T_1} = 551 \text{ mL} \times \frac{373 \text{ K}}{298 \text{ K}} = 690 \text{ mL}$$

(4.90 × 10² mL to 3 s.f.)

Gay-Lussac's Law: A sample of CH_4 gas in a sealed flask is initially at a pressure of 37.5 kPa and a temperature of -150°C. What will the pressure be after warming the flask to 25°C?

$$\frac{P_1}{T_1} = \frac{P_2}{T_2} \Rightarrow P_2 = P_1 \times \frac{T_2}{T_1} = 37.5 \text{ kPa} \times \frac{298 \text{ K}}{77 \text{ K}} = 145 \text{ kPa}$$

(technically only 2 s.f. from 77 K, though)

liquid nitrogen bath!

Combined Gas Law: A sample of N_2 gas initially occupies 7.11 cm³ in a sealed syringe at a pressure of 1.25 atm and a temperature of 25°C. If we push down on the syringe, compressing the gas to a volume of only 5.00 cm³, the temperature increases slightly to 30°C. What is the new pressure after compression?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \Rightarrow P_2 = P_1 \times \frac{V_1}{V_2} \times \frac{T_2}{T_1} = 1.25 \text{ atm} \times \frac{7.11 \text{ cm}^3}{5.00 \text{ cm}^3} \times \frac{303 \text{ K}}{298 \text{ K}} = 1.84 \text{ atm}$$

relatively large increase in P due to decrease in volume small increase in P due to change in temp.

The Ideal Gas Law combines Boyle's, Charles', Gay-Lussac's and Avogadro's Law (next page) into a single expression relating P, V, T and the number of moles (n):

The Ideal Gas Law:

$$PV = nRT$$

The ideal gas constant, R, is equal to 8.314 kPa L mol⁻¹ K⁻¹ or 8.314 J mol⁻¹ K⁻¹.

Example 1: What volume, in L, is occupied by 1.00 mole of an ideal gas at 1.00 × 10⁵ Pa and 0°C?

$$V = \frac{nRT}{P} = \frac{(1.00 \text{ mol})(8.314 \text{ kPa} \cdot \text{L} \cdot \text{mol}^{-1} \cdot \text{K}^{-1})(273 \text{ K})}{100 \text{ kPa}} = 22.7 \text{ L} \quad (\text{see next page...})$$

Example 2: What is the mass, in grams, of a sample of nitrogen gas at a pressure of 511 kPa, a volume of 25.0 L, and a temperature of 25°C?

$$n = \frac{PV}{RT} = \frac{(511 \text{ kPa})(25.0 \text{ L})}{(8.314 \text{ kPa} \cdot \text{L} \cdot \text{mol}^{-1} \cdot \text{K}^{-1})(298 \text{ K})} = 5.156 \dots \text{ mol } N_2$$

$$5.156 \dots \text{ mol } N_2 \times \frac{28.02 \text{ g}}{1 \text{ mol}} = 144 \text{ g}$$

Ib Chemistry Hl Stoichiometry

Joacim Rocklöv



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