

The Hydrogen Atom – Student Guide

Background Material

Carefully read the background pages created through Eureka, Slides, and Experiments and answer the following questions in detail and with understanding.

Question 1: Complete the following table which compares how the Bohr model and the Quantum model compare the Hydrogen atom. In each case, they both describe things in a unique way and in some ways they do not.

Concept Name	Bohr Model	Quantum Model
Electrons	positively charged particles in the atom	negatively charged particles in the atom
Orbitals	negatively charged and orbit the nucleus	negatively charged clouds
Location of the electron in the atom without	is in two fixed paths	electron cloud pattern
Energy of the electron in the atom without	always has the same energy	varies the energy level
Energy of the electron in the atom without	$E = -13.6 \text{ eV}$	$E = -13.6 \text{ eV}$

Question 2: Complete the following table which shows the differences of two different models by putting predictions about the spectrum of hydrogen in the table. The table is completed by the 19th and 20th century's scientific knowledge, another hypothesis, another energy, and the same method through which we will observe.

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J Elliott



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