

QUESTION 10

The three test tubes, X, Y and Z below, contain colourless solutions. The solutions in each test tube can either be potassium iodide or sodium carbonate or magnesium sulphate. Tests are carried out to determine which test tube contains which salts.



The following tests were carried out and the observations were made:

- Barium chloride solution is added to solution X; the solution remains colourless.
- Barium chloride solution is added to Y and Z; a white precipitate is formed in both cases.
- Dilute nitric acid is added to the precipitate in solution Y; the precipitate dissolves.

(2)

10.1 Define the term *dissociation*.

10.2 Give a reason why barium chloride solution can be used as electrolyte.

(2)

10.3 Write down the name of the precipitate in test tube Z.

(1)

Use the above information to identify the solutions in each of the test tubes X, Y and Z.

10.4 Write down the LETTER that represents the test tube and next to it, the NAME of the solution and the reason for the choice.

(6)

10.5 What type of reaction takes place between the precipitate in Y and the nitric acid?

(1)

[12]

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Michael Rajnik

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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Table of Contents Grade Exemplar Physical Sciences P2

1. Understanding the eBook Grade Exemplar Physical Sciences P2
 - The Rise of Digital Reading Grade Exemplar Physical Sciences P2
 - Advantages of eBooks Over Traditional Books
2. Identifying Grade Exemplar Physical Sciences P2
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade Exemplar Physical Sciences P2
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grade Exemplar Physical Sciences P2
 - Personalized Recommendations

- Grade Exemplar Physical Sciences P2 User Reviews and Ratings
- Grade Exemplar Physical Sciences P2 and Bestseller Lists
- 5. Accessing Grade Exemplar Physical Sciences P2 Free and Paid eBooks
 - Grade Exemplar Physical Sciences P2 Public Domain eBooks
 - Grade Exemplar Physical Sciences P2 eBook Subscription Services
 - Grade Exemplar Physical Sciences P2 Budget-Friendly Options
- 6. Navigating Grade Exemplar Physical Sciences P2 eBook Formats
 - ePub, PDF, MOBI, and More
 - Grade Exemplar Physical Sciences P2 Compatibility with Devices
 - Grade Exemplar Physical Sciences P2 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Grade Exemplar Physical Sciences P2
 - Highlighting and Note-Taking Grade Exemplar Physical Sciences P2
 - Interactive Elements Grade Exemplar Physical Sciences P2
- 8. Staying Engaged with Grade Exemplar Physical Sciences P2
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade Exemplar Physical Sciences P2
- 9. Balancing eBooks and Physical Books Grade Exemplar Physical Sciences P2
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grade Exemplar Physical Sciences P2
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Grade Exemplar Physical Sciences P2
 - Setting Reading Goals Grade Exemplar Physical Sciences P2
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grade Exemplar Physical Sciences P2
 - Fact-Checking eBook Content of Grade Exemplar Physical Sciences P2

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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