

1. What amount of oxygen, O_2 , (in moles) contains 1.8×10^{22} molecules?

- | | | | |
|-------------------------|-------|-------------------------|--------|
| ☐ A | 0.030 | ☐ B | 3.0 |
| ☐ C | 0.30 | ☐ D | 0.0030 |

2. Which compound has the empirical formula with the greatest mass?

- | | | | |
|-------------------------|-------------|-------------------------|-------------|
| ☐ A | C_6H_6 | ☐ B | C_5H_{10} |
| ☐ C | C_4H_{10} | ☐ D | C_2H_6 |

3. $__C2H2(g) + __O2(g) \rightarrow __CO2(g) + __H2O(g)$

When the equation above is balanced, what is the coefficient for oxygen?

- | | | | |
|-------------------------|---|-------------------------|---|
| ☐ A | 2 | ☐ B | 3 |
| ☐ C | 5 | ☐ D | 4 |

4. 3.0 dm³ of sulfur dioxide is reacted with 2.0 dm³ of oxygen according to the equation below.



What volume of sulfur trioxide (in dm³) is formed? (Assume the reaction goes to completion and all gases are measured at the same temperature and pressure.)

- | | | | |
|-------------------------|-----|-------------------------|-----|
| ☐ A | 5.0 | ☐ B | 3.0 |
| ☐ C | 2.0 | ☐ D | 4.0 |

5. What will happen to the volume of a fixed mass of gas when its pressure and temperature (in Kelvin) are both doubled?

- | | | | |
|-------------------------|---------------------|-------------------------|---------------------------------|
| ☐ A | It will not change. | ☐ B | The change cannot be predicted. |
| ☐ C | It will increase. | ☐ D | It will decrease. |

6. What amount (in moles) is present in 2.0 g of sodium hydroxide, NaOH?

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