

Name: \_\_\_\_\_

Date: \_\_\_\_\_

## Student Exploration: Ionic Bonds

**Vocabulary:** chemical family, electron affinity, ion, ionic bond, metal, nonmetal, octet rule, shell, valence electron

**Prior Knowledge Questions** (Do these BEFORE using the Gizmo.)

1. Nate and Clara are drawing pictures with markers. There are 8 markers in a set. Nate has 9 markers and Clara has 7. What can Nate and Clara do so that each of them has a full set?

\_\_\_\_\_ Nate can share a marker with Clara, so they each a full set 8 markers \_\_\_\_\_

2. Maggie is sitting at a table with Fred and Florence. Maggie has 10 markers, but Fred and Florence each have only 7 markers. How can they share markers so each has 8?

\_\_\_\_\_ Maggie can give Fred and Florence each 1 marker, which would leave everybody with 9 markers. \_\_\_\_\_

### Gizmo Warm-up

Just like students sharing markers, atoms sometimes share or swap electrons. By doing this, atoms form bonds. The *Ionic Bonds* Gizmo™ allows you to explore how **ionic bonds** form.

To begin, check that **Sodium (Na)** and **Chlorine (Cl)** are selected from the menus at right. Click **Play** (▶) to see electrons orbiting the nucleus of each atom. (Note: These atom models are simplified and not meant to be realistic.)



1. Each atom consists of a central nucleus and several **shells** that contain electrons. The outermost electrons are called **valence electrons**. (Inner electrons are not shown.)

How many valence electrons does each atom have? Sodium: 1 Chlorine: 7

2. Click **Pause** (⏏). Elements can be classified as **metals** and **nonmetals**. Metals do not hold on to their valence electrons very tightly, while nonmetals hold their electrons tightly. **Electron affinity** is a measure of how tightly the valence electrons are held.

# Exploration Guide Ionic Bonds Answer Key

**A Gutmann**



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## **Table of Contents Exploration Guide Ionic Bonds Answer Key**

1. Understanding the eBook Exploration Guide Ionic Bonds Answer Key
  - The Rise of Digital Reading Exploration Guide Ionic Bonds Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Exploration Guide Ionic Bonds Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Exploration Guide Ionic Bonds Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Exploration Guide Ionic Bonds Answer Key
  - Personalized Recommendations
  - Exploration Guide Ionic Bonds Answer Key User Reviews and Ratings
  - Exploration Guide Ionic Bonds Answer Key and Bestseller Lists
5. Accessing Exploration Guide Ionic Bonds Answer Key Free and Paid eBooks
  - Exploration Guide Ionic Bonds Answer Key Public Domain eBooks
  - Exploration Guide Ionic Bonds Answer Key eBook Subscription Services

- Exploration Guide Ionic Bonds Answer Key Budget-Friendly Options
- 6. Navigating Exploration Guide Ionic Bonds Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Exploration Guide Ionic Bonds Answer Key Compatibility with Devices
  - Exploration Guide Ionic Bonds Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Exploration Guide Ionic Bonds Answer Key
  - Highlighting and Note-Taking Exploration Guide Ionic Bonds Answer Key
  - Interactive Elements Exploration Guide Ionic Bonds Answer Key
- 8. Staying Engaged with Exploration Guide Ionic Bonds Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Exploration Guide Ionic Bonds Answer Key
- 9. Balancing eBooks and Physical Books Exploration Guide Ionic Bonds Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Exploration Guide Ionic Bonds Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Exploration Guide Ionic Bonds Answer Key
  - Setting Reading Goals Exploration Guide Ionic Bonds Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Exploration Guide Ionic Bonds Answer Key
  - Fact-Checking eBook Content of Exploration Guide Ionic Bonds Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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