

Step 3 From the number of moles of Fe^{3+} and SCN^- initially present in each mixture, and the number of moles of Fe^{3+} and SCN^- used up in forming FeSCN^{2+} , calculate the number of moles of Fe^{3+} and SCN^- that remain in each mixture at equilibrium. Use Equation 4. Enter the results in columns 3 and 4 of the table.

Step 4 Use Equation 3 and the results of Step 3 to find the concentrations of all of the species at equilibrium. The volume of the mixture is 10.00 mL, or 0.0100 liter, in all cases. Enter the values in columns 5 and 7 of the table.

Step 5 Calculate K_c for the reaction for each of the mixtures by substituting values for the equilibrium concentrations of Fe^{3+} , SCN^- , and FeSCN^{2+} in Equation 2.

Step 6 Calculate the mean value for K_c and the standard deviation. (See Appendix VIII.)

Mixture	Initial Moles		Equilibrium Moles			Equilibrium Concentration			K_c
	Fe^{3+}	SCN^-	Fe^{3+}	SCN^-	FeSCN^{2+}	$[\text{Fe}^{3+}]$	$[\text{SCN}^-]$	$[\text{FeSCN}^{2+}]$	
1	$\text{---} \times 10^{-4}$	$\text{---} \times 10^{-4}$	$\text{---} \times 10^{-4}$	$\text{---} \times 10^{-4}$	$\text{---} \times 10^{-4}$	$\text{---} \times 10^{-4} \text{ M}$	$\text{---} \times 10^{-4} \text{ M}$	$\text{---} \times 10^{-4} \text{ M}$	---
2	---	---	---	---	---	---	---	---	---
3	---	---	---	---	---	---	---	---	---
4	---	---	---	---	---	---	---	---	---
5	---	---	---	---	---	---	---	---	---

Mean value of K_c --- Standard deviation ---

(continued on following page)

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