

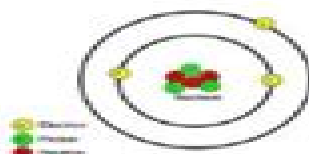
## Atomic Structure – Study Guide

### Grade 8 – Physical Science

**Atom** -- smallest particle of an element that has all the properties of that element.

**3 main subatomic particles:**

1. protons
2. neutrons
3. electrons



**nucleus** -- core of the atom, it is positively charged 99.9 % of the atoms mass is in the nucleus. Protons and neutrons are found in the **nucleus**.

**1. Protons** -- have a positive (+) charge. Found in the nucleus of the atom.

The **number of protons** determines what element an atom is.

**Example:** Carbon has 6 protons

**Atomic Number** = the **number of protons**, each element has it own unique number of protons.

**2. Neutrons** -- a neutral charge (no charge) found in the nucleus. These determine whether an element is radioactive.

Protons and neutrons are about *one atomic mass unit (amu)*. Electrons have a much smaller mass -- it takes almost 2000 electrons to equal 1 amu.

**Atomic Mass** = the total number of protons and neutrons.

**Mass Number** = Atomic Mass that is rounded.

To find just how many **neutrons** an atom has:

# neutrons = atomic mass – atomic #

**Isotopes** - atoms of the same element that have different number of neutrons.

**Carbon-12**

6 protons

6 neutrons

**Carbon-14**

6 protons

8 neutrons

**NOTICE** – the number of **protons DOES NOT** change if it is to be carbon -- only the number of neutrons.

**3. Electrons** - have a negative (-) charge. Found moving around the atom is a space called orbital's (shells).

Those electrons farthest away from the nucleus are called **valence electrons**, and are involved in the formation of chemical bonds.

**In an uncharged atom the number of protons is equal to the number of electrons.**

**ion** - when an atom gains or loses an *electron*. (has a charge)

If an atom gains electrons = negative (-) charge

If an atom loses electrons = positive (+) charge

**How to find:**

# protons = atomic #

# neutrons = atomic mass – atomic #

# electrons = atomic # – charge (if any)

# Matter And Atomic Structure Study Guide Answers

**Jianjun Gao**



## **Matter And Atomic Structure Study Guide Answers:**

## Reviewing **Matter And Atomic Structure Study Guide Answers**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Matter And Atomic Structure Study Guide Answers**," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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