# Limiting Reagent & Percent Yield Practice Worksheet

- When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.
  - a. Write the balanced equation for the reaction given above:

 $CuCl_2 + NaNO_3 \rightarrow Cu(NO_3)_2 + NaCl$  $CuCl_2 + 2 NaNO_3 \rightarrow Cu(NO_3)_2 + 2 NaCl$ 

b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

15 g CuCl<sub>2</sub> x 1 mol CuCl<sub>2</sub> x 2 mol NaCl x 58.443 g NaCl = 13 g NaCl 134.362g CuCl<sub>2</sub> 1 mol CuCl<sub>2</sub> 1 mol NaCl

20 g NaNO<sub>3</sub> x <u>1 mol NaNO<sub>3</sub></u> x <u>2 mol NaCl</u> x <u>58.443 g NaCl</u> = 13.8 g NaCl 84.997g NaNO<sub>3</sub> 2 mol NaNO<sub>3</sub> 1 mol NaCl

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? copper (11) chloride
- d. How many grams of copper (II) nitrate is formed?

15g CuCl<sub>2</sub> x 1 mol CuCl<sub>2</sub> x 1 mol Cu(NO<sub>3</sub>)<sub>2</sub> x 187.467 g Cu(NO<sub>3</sub>)<sub>2</sub> = 20.9 g Cu(NO<sub>3</sub>)<sub>2</sub> 134.452g CuCl<sub>2</sub> 1 mol CuCl<sub>2</sub> 1 mol Cu(NO<sub>3</sub>)<sub>2</sub>

e. How much of the excess reagent is left over in this reaction?

15g CuCl<sub>2</sub> x 1 mol CuCl<sub>2</sub> x 2 mol NaNO<sub>3</sub> x 84,997g NaNO<sub>3</sub> = 19 g NaNO<sub>3</sub> 134,452g CuCl<sub>2</sub> 1 mol CuCl<sub>2</sub> 1 mol NaNO<sub>3</sub>

20 g NaNO<sub>5</sub> - 19 g NaNO<sub>5</sub> = 1 g NaNO<sub>5</sub>

f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction? Percent yield = actual yield / theoretical yield × 100 percent

11.3/13.0 x 100% = 86.9%

- When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.
  - a. Balance the following equation:

 $Pb(NO_3)_{2 \text{ (sq)}} + NaI_{\text{ (sq)}} \rightarrow PbI_{2 \text{ (s)}} + NaNO_{3 \text{ (sq)}}$  $Pb(NO_3)_{2 \text{ (sq)}} + 2 NaI_{\text{ (sq)}} \rightarrow PbI_{2 \text{ (s)}} + 2 NaNO_{3 \text{ (sq)}}$ 

b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

25.0 g Pb(NO<sub>3</sub>)<sub>2</sub> x 1 mol Pb(NO<sub>3</sub>)<sub>2</sub> x 2 mol NaNO<sub>3</sub> x 84.997g NaNO<sub>3</sub> = 12.8 g NaNO<sub>3</sub> 331.214 g Pb(NO<sub>3</sub>)<sub>2</sub> 1 mol Pb(NO<sub>3</sub>)<sub>2</sub> 1 mol NaNO<sub>3</sub>

15.0 g Nal x 1 mol Nal x 2 mol NaNO<sub>3</sub> x 84.997g NaNO<sub>3</sub> = 8.51 g NaNO<sub>3</sub> 149.89g Nal 2 mol Nal 1 mol NaNO<sub>3</sub>

8.51 g NaNO3 can be made

# **Limiting Reagent And Percent Yield Key**

Shasha Hu

#### **Limiting Reagent And Percent Yield Key:**

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