

Periodic Trends Worksheet

Directions: Use the table to determine the following properties:

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

K, Na, O, S, Cl

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

K, Na, S, O, Cl

3. Which atom has the highest first ionization energy than sodium?

Chlorine (because it has the most electronegativity and is closest to sodium)

4. Which two elements in the same family generally have similar properties?

There have been two atoms that follow a strong periodic trend: potassium and sodium

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (increases to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the lowest?

Period 7, francium (because it's the furthest from the nucleus)

7. Which trend in ionization energy atoms across a period and the periodic table? Which element has the highest?

Increases across the periodic table, fluorine has the highest

8. Write the atoms in each group that have the largest atomic radius.

Li	Be
Na	Mg
K	Ca
Rb	Sr
Cs	Ba
Fr	Ra

Al	Si	P
Ge	As	Se
Sb	Te	I
Bi	Po	At

9. Circle the atom in each group that has the greatest electronegativity.

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Yijin Wang

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the Yijin Wang text bar.

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