

## 15 Hydrocarbons

### Alkanes

#### Chemical Reactivity

Alkanes do not react with many reagents.

This is because the C-C bond and the C-H bond are relatively strong and so cannot form stronger bonds with other elements that easily and there are no polar side groups.

Their main reactions are combustion reactions in air although they can react with halogens in the presence of UV light.

#### Combustion

##### Complete Combustion

In excess oxygen alkanes will burn with complete combustion.

The products of **complete** combustion are  $\text{CO}_2$  and  $\text{H}_2\text{O}$ .



Alkanes readily burn in the presence of oxygen. This combustion of alkanes is highly exothermic, explaining their use as **fuels**.

#### Reaction of alkanes with bromine / chlorine in UV light

##### Synthesis of chloroalkanes

In the presence of **UV light** alkanes react with chlorine to form a mixture of products with the halogens substituting hydrogen atoms.

##### Overall Reaction



This is the overall reaction, but a more complex mixture of products is actually formed.

To understand this reaction fully we must look in detail at how it proceeds step by step. This is called its **mechanism**.

The **MECHANISM** for this reaction is called a **FREE RADICAL SUBSTITUTION**.

It proceeds via a series of steps:  
**STEP ONE: Initiation**  
**STEP TWO: Propagation**  
**STEP THREE: Termination**

##### STEP ONE Initiation

Essential condition: UV light



The UV light supplies the energy to break the Cl-Cl bond. It is broken in preference to the others as it is the weakest.

The bond has broken in a process called **homolytic fission**.

**each atom gets one electron from the covalent bond**

When a bond breaks by homolytic fission it forms **Free Radicals**.

**Free Radicals** do not have a charge and are represented by a  $\bullet$

##### DEFINITION

A **Free Radical** is a reactive species which possess an **unpaired electron**.

##### STEP TWO Propagation



The chlorine free radicals are very reactive and remove an H from the methane leaving a methyl free radical.

The methyl free radical reacts with a  $\text{Cl}_2$  molecule to produce the main product and another Cl free radical.

All propagation steps have a **free radical** in the **reactants** and in the **products**.

As the Cl free radical is regenerated, it can react with several more alkane molecules in a **CHAIN REACTION**.

# Hydrocarbon Compounds Study Guide Answers

**Clyde Metz**



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## **Table of Contents Hydrocarbon Compounds Study Guide Answers**

1. Understanding the eBook Hydrocarbon Compounds Study Guide Answers
  - The Rise of Digital Reading Hydrocarbon Compounds Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Hydrocarbon Compounds Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Hydrocarbon Compounds Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Hydrocarbon Compounds Study Guide Answers
  - Personalized Recommendations
  - Hydrocarbon Compounds Study Guide Answers User Reviews and Ratings
  - Hydrocarbon Compounds Study Guide Answers and Bestseller Lists
5. Accessing Hydrocarbon Compounds Study Guide Answers Free and Paid eBooks
  - Hydrocarbon Compounds Study Guide Answers Public Domain eBooks
  - Hydrocarbon Compounds Study Guide Answers eBook Subscription Services
  - Hydrocarbon Compounds Study Guide Answers Budget-Friendly Options
6. Navigating Hydrocarbon Compounds Study Guide Answers eBook Formats

- ePub, PDF, MOBI, and More
  - Hydrocarbon Compounds Study Guide Answers Compatibility with Devices
  - Hydrocarbon Compounds Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Hydrocarbon Compounds Study Guide Answers
    - Highlighting and Note-Taking Hydrocarbon Compounds Study Guide Answers
    - Interactive Elements Hydrocarbon Compounds Study Guide Answers
  8. Staying Engaged with Hydrocarbon Compounds Study Guide Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Hydrocarbon Compounds Study Guide Answers
  9. Balancing eBooks and Physical Books Hydrocarbon Compounds Study Guide Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Hydrocarbon Compounds Study Guide Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Hydrocarbon Compounds Study Guide Answers
    - Setting Reading Goals Hydrocarbon Compounds Study Guide Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Hydrocarbon Compounds Study Guide Answers
    - Fact-Checking eBook Content of Hydrocarbon Compounds Study Guide Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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