

Chapter 1

Vector Analysis

Problem 1.1

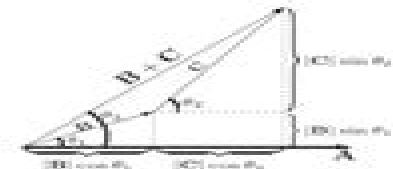
- (a) From the diagram, $|\mathbf{B} + \mathbf{C}| \cos \theta_3 = |\mathbf{B}| \cos \theta_1 + |\mathbf{C}| \cos \theta_2$. Multiply by $|\mathbf{A}|$.
 $|\mathbf{A}||\mathbf{B} + \mathbf{C}| \cos \theta_3 = |\mathbf{A}||\mathbf{B}| \cos \theta_1 + |\mathbf{A}||\mathbf{C}| \cos \theta_2$.

So: $\mathbf{A} \cdot (\mathbf{B} + \mathbf{C}) = \mathbf{A} \cdot \mathbf{B} + \mathbf{A} \cdot \mathbf{C}$. (Dot product is distributive)

Similarly: $|\mathbf{B} + \mathbf{C}| \sin \theta_3 = |\mathbf{B}| \sin \theta_1 + |\mathbf{C}| \sin \theta_2$. Multiply by $|\mathbf{A}| \hat{n}$.
 $|\mathbf{A}||\mathbf{B} + \mathbf{C}| \sin \theta_3 \hat{n} = |\mathbf{A}||\mathbf{B}| \sin \theta_1 \hat{n} + |\mathbf{A}||\mathbf{C}| \sin \theta_2 \hat{n}$.

If $\hat{n}$ is the unit vector pointing out of the page, it follows that

$\mathbf{A} \times (\mathbf{B} + \mathbf{C}) = (\mathbf{A} \times \mathbf{B}) + (\mathbf{A} \times \mathbf{C})$. (Cross product is distributive)



- (b) For the general case, see G. E. Hay's *Vector and Tensor Analysis*, Chapter 1, Section 7 (dot product) and Section 8 (cross product)

Problem 1.2

The triple cross-product is *not* in general associative. For example, suppose $\mathbf{A} = \mathbf{B}$ and $\mathbf{C}$ is perpendicular to $\mathbf{A}$, as in the diagram. Then $(\mathbf{B} \times \mathbf{C})$ points out-of-the-page, and $\mathbf{A} \times (\mathbf{B} \times \mathbf{C})$ points down, and has magnitude ABC . But $(\mathbf{A} \times \mathbf{B}) = \mathbf{0}$, so $(\mathbf{A} \times \mathbf{B}) \times \mathbf{C} = \mathbf{0} \neq \mathbf{A} \times (\mathbf{B} \times \mathbf{C})$.

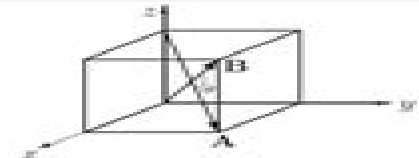


Problem 1.3

$$\mathbf{A} = +1\hat{x} + 1\hat{y} - 1\hat{z}; A = \sqrt{3}; \mathbf{B} = 1\hat{x} + 1\hat{y} + 1\hat{z}; B = \sqrt{3}.$$

$$\mathbf{A} \cdot \mathbf{B} = +1 + 1 - 1 = 1 = AB \cos \theta = \sqrt{3}\sqrt{3} \cos \theta \Rightarrow \cos \theta = \frac{1}{3}.$$

$$\theta = \cos^{-1} \left(\frac{1}{3} \right) \approx 70.5288^\circ$$



Problem 1.4

The cross-product of any two vectors in the plane will give a vector perpendicular to the plane. For example, we might pick the base ($\mathbf{A}$) and the left side ($\mathbf{B}$):

$$\mathbf{A} = -1\hat{x} + 2\hat{y} + 0\hat{z}; \mathbf{B} = -1\hat{x} + 0\hat{y} + 3\hat{z}.$$

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web in 1916 american chemist gilbert n lewis introduced bond lines to electron dot structures these structures also known as lewis structures or electron dot structures are drawings that visually demonstrate how electrons are shared and arranged around atoms the electrons denoted as dots are called lone pairs and belong to an individual

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web lewis dot structures are useful for describing chemical bonds but have some flaws a lewis structure is a picture of a molecule that shows the covalent bonds and pairs of free electrons the octet rule is the basis for lewis structures

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web lewis dot formula also called an electron dot formula shows the valence electrons indicating the bonding between atoms the following guidelines will help draw the electron dot formulas correctly

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web lewis structures extend the concept of the electron dot diagram by adding lines between atoms to represent shared pairs in a chemical bond lewis structures show each atom and its position in the structure of the molecule using its chemical symbol

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web lewis dot structures also called electron dot structures are diagrams that describe the chemical bonding between atoms in a molecule they also display the total number of lone pairs present in each of the atoms that constitute the molecule

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web oct 29 2021 contributors and attributions learning objectives to draw lewis structures for molecules and polyatomic ions with one central atom introduction to lewis structures a lewis structure is a way to show how atoms share electrons when they form a molecule lewis structures show all of the valence electrons in an atom or molecule

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web lewis structures we also use lewis symbols to indicate the formation of covalent bonds which are shown in lewis structures drawings that describe the bonding in molecules and polyatomic ions for example when two chlorine atoms form a chlorine molecule they share one pair of electrons

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web this online quiz is intended to give you extra practice in identifying and drawing lewis dot structures as well as predicting ion formation this quiz aligns with the following ngss standard s hs ps1 1

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web jan 30 2023 a lewis structure is a very simplified representation of the valence shell electrons in a molecule it is used to show how the electrons are arranged around individual atoms in a molecule electrons are shown as dots or for bonding electrons as a line between the two atoms

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web jan 30 2023 follow these simple steps to draw lewis dot structures draw the atoms on paper and put dots around them to represent valence electrons of the atom be sure to have the correct number of electrons if the species is an ion add or subtract electrons corresponding to the charge of the ion

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