UNIT REVIEW - BONDING

Intramolecular Forces

- 1. Define electronegativity
 - the tendency of atoms to attract e-
- as you go across the periodic table, and <u>cleareases</u> as you go 2. Electronegativity (ncrocuso) down a group.
- 3. Explain why atoms with smaller atomic radii have stronger electronegativities

 - smaller atoms means e's are closer to the nucleus, which has the protons if you decrease the distance between protons + e's, you increase attractive forces.
- 4. Define intramolecular force
 - forces that hold atoms together: ionic, covalent, polar covalent bond
- Compare and contrast covalent, polar covalent, and ionic bond.

Covalent	Polar covalent	Ionic
- e- are shared equally - between a non-motali - difference in electronegativity is ±0.2 - otoms are newbal	- e- are shared unequally - between a non-metals - difference in electronegotisty is 0.3-1.6 - atoms are partially charged	- e are transfered - between metal + non-meta - difference in electronegativity is ≥1.7 - atoms have a charge

6. What is a partial charge and how is it different from an ionic charge?

A partial charge results from an unequal shaving of elections between a atoms An ionic change results from a transfer of elections from I clom to another

Flocabulary Transformation Answer Key Unit

Raffaela Di Napoli

Flocabulary Transformation Answer Key Unit:

Reviewing Flocabulary Transformation Answer Key Unit: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "Flocabulary Transformation Answer Key Unit," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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