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Limiting Reagent Worksheet

1) *When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.*

a) Write the balanced equation for the reaction given above:



b) If 45 grams of copper (II) chloride react with 25 grams of sodium nitrate, how much sodium chloride can be formed? **17 g.**

c) What is the limiting reagent for the reaction? **NaNO₃**

d) How many grams of copper(II) nitrate is formed? **28 g.**

e) How much of the excess reagent is left over in this reaction? **25 g.**

f) If 11.3 grams of sodium chloride are formed in the reaction, what is the percent yield of this reaction? **66%**

2) *When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.*

a) Balance the following equation:



b) If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed? **8.51 g.**

c) What is the limiting reagent in the reaction? **NaI**

d) How many grams of lead(II) iodide is formed? **23.1 g.**

e) How much of the nonlimiting reagent will be left over from the reaction? **8.4 g.**

f) If 6.0 grams of sodium nitrate are formed in the reaction, what is the percent yield of this reaction? **71%**

3) *1000 grams of sodium chloride is combined with 2000 grams of barium phosphate.*

a) Balance the following equation:



b) What is the limiting reactant? **NaCl**

c) How many grams of excess reactant are left? **283 g.**

d) How many grams of BaCl₂ are produced? **1782 g.**

e) How many grams of Na₃PO₄ are formed? **935.0 g.**

f) If 1,483 grams of BaCl₂ are formed in the reaction, what is the percent yield of this reaction?
83.22%

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Flinn Scientific Limiting Reagent Lab Answers

David Kirk



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Table of Contents Flinn Scientific Limiting Reagent Lab Answers

1. Understanding the eBook Flinn Scientific Limiting Reagent Lab Answers
 - The Rise of Digital Reading Flinn Scientific Limiting Reagent Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Flinn Scientific Limiting Reagent Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Flinn Scientific Limiting Reagent Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Flinn Scientific Limiting Reagent Lab Answers
 - Personalized Recommendations
 - Flinn Scientific Limiting Reagent Lab Answers User Reviews and Ratings
 - Flinn Scientific Limiting Reagent Lab Answers and Bestseller Lists

5. Accessing Flinn Scientific Limiting Reagent Lab Answers Free and Paid eBooks
 - Flinn Scientific Limiting Reagent Lab Answers Public Domain eBooks
 - Flinn Scientific Limiting Reagent Lab Answers eBook Subscription Services
 - Flinn Scientific Limiting Reagent Lab Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Flinn Scientific Limiting Reagent Lab Answers Compatibility with Devices
 - Flinn Scientific Limiting Reagent Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Flinn Scientific Limiting Reagent Lab Answers
 - Highlighting and Note-Taking Flinn Scientific Limiting Reagent Lab Answers
 - Interactive Elements Flinn Scientific Limiting Reagent Lab Answers
8. Staying Engaged with Flinn Scientific Limiting Reagent Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Flinn Scientific Limiting Reagent Lab Answers
9. Balancing eBooks and Physical Books Flinn Scientific Limiting Reagent Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Flinn Scientific Limiting Reagent Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Flinn Scientific Limiting Reagent Lab Answers
 - Setting Reading Goals Flinn Scientific Limiting Reagent Lab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Flinn Scientific Limiting Reagent Lab Answers
 - Fact-Checking eBook Content of Flinn Scientific Limiting Reagent Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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