

Name: _____

Hour: _____

Lab 7: Flame Test Lab

Prelab Reading

The normal electron configuration of atoms or ions of an element is known as the "ground state". In this most stable energy state all of the electrons are in the lowest energy levels available. When atoms or ions in the ground state are heated to high temperatures some electrons may absorb energy and "jump" to a higher energy level. The element is then said to be in an "excited state". This excited configuration is unstable and the electrons will "fall back" to their normal positions of lower energy. As the electrons return to their normal levels the energy that was absorbed is now emitted in the form of electromagnetic radiation. Some of this radiation is in the form of visible light.

The color of this light can be used as a means of identifying the elements involved. This analysis is known as a flame test. Only metals, with their loosely held electrons, are excited by the flame of a bunsen burner. Thus flame tests are useful in the identification of metallic ions. Many metallic ions exhibit characteristic colors when vaporized by the burner flame. In this experiment characteristic colors of several different metallic ions will be observed and unknown ions will be identified by means of the flame test.

Before beginning the pre-lab questions below, be sure to read the procedure on the next page.

Pre-Lab Questions:

1. What is the difference between ground state and an excited state?

Ground state: all electrons are at the lowest possible energy levels.

Excited state: Atoms/ions absorb energy and go to a higher energy.

2. What does the word "emit" mean?

Emit is defined as being produced (i.e. the energy that was absorbed is produced in the form of electromagnetic radiation).

3. In this experiment, where are the atoms getting their excess energy from?

The electrons are getting the excess energy from a bunsen burner.

4. Why do different atoms emit different colors of light?

Atoms emit different colors because of their chemical makeup.

5. Why is it necessary to clean the nichrome wires between each flame test?

It is necessary to clean the wires between the flame tests because if you don't, the atoms might not display their characteristic color because they were mixed with another atom.

Flame Test Lab Answers

Lauren Gardner



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Table of Contents Flame Test Lab Answers

1. Understanding the eBook Flame Test Lab Answers
 - The Rise of Digital Reading Flame Test Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Flame Test Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Flame Test Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Flame Test Lab Answers
 - Personalized Recommendations
 - Flame Test Lab Answers User Reviews and Ratings
 - Flame Test Lab Answers and Bestseller Lists
5. Accessing Flame Test Lab Answers Free and Paid eBooks
 - Flame Test Lab Answers Public Domain eBooks
 - Flame Test Lab Answers eBook Subscription Services
 - Flame Test Lab Answers Budget-Friendly Options
6. Navigating Flame Test Lab Answers eBook Formats

- ePub, PDF, MOBI, and More
- Flame Test Lab Answers Compatibility with Devices
- Flame Test Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Flame Test Lab Answers
 - Highlighting and Note-Taking Flame Test Lab Answers
 - Interactive Elements Flame Test Lab Answers
- 8. Staying Engaged with Flame Test Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Flame Test Lab Answers
- 9. Balancing eBooks and Physical Books Flame Test Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Flame Test Lab Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Flame Test Lab Answers
 - Setting Reading Goals Flame Test Lab Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Flame Test Lab Answers
 - Fact-Checking eBook Content of Flame Test Lab Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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