

$$M_{\text{AgNO}_3} = \frac{n_{\text{AgNO}_3}}{V_{\text{AgNO}_3}} = \frac{25.513 \text{ g} + 149.87 \frac{\text{g}}{\text{mol}}}{100.00 \text{ mL}} + \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{1.569 \frac{\text{mol}}{\text{L}}}$$

4. Confirm that the molarity is the same molarity that you calculated in step 3.

Yes, the molarity is the same as the molarity calculated in step 3.

Part B: Unknown solutions

Using a method similar to the one used in part A, create and carry out a procedure for determining the molarities of the unknown AgNO_3 solution and the unknown NaCl solution. Round your data and show your calculations.

Unknown AgNO_3

Upon the addition of solid NaCl , solid AgCl was formed.

AgCl formed (g) = 2.32019 g

$$\text{mass}(\text{AgNO}_3) = \frac{M_r(\text{AgNO}_3)}{M_r(\text{AgCl})} \times \text{mass}(\text{AgCl}) = \frac{149.87}{143.32} \frac{\text{g}}{\text{mol}} \times 2.32019 \text{ g} = \boxed{2.3500 \text{ g}}$$

$$M_{\text{AgNO}_3} = \frac{n_{\text{AgNO}_3}}{V_{\text{AgNO}_3}} = \frac{2.3500 \text{ g} + 149.87 \frac{\text{g}}{\text{mol}}}{100.00 \text{ mL}} + \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{0.16189 \frac{\text{mol}}{\text{L}}}$$

Unknown NaCl

Upon the addition of solid AgNO_3 , solid AgCl was formed.

AgCl formed (g) = 10.7907 g

$$\text{mass}(\text{NaCl}) = \frac{M_r(\text{NaCl})}{M_r(\text{AgCl})} \times \text{mass}(\text{AgCl}) = \frac{58.443}{143.32} \frac{\text{g}}{\text{mol}} \times 10.7907 \text{ g} = \boxed{4.4002 \text{ g}}$$

$$M_{\text{NaCl}} = \frac{n_{\text{NaCl}}}{V_{\text{NaCl}}} = \frac{4.4002 \text{ g} + 58.443 \frac{\text{g}}{\text{mol}}}{100.00 \text{ mL}} + \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{0.75294 \frac{\text{mol}}{\text{L}}}$$

Conclusion questions

1. If you have a company that is using this method to create AgCl , knowing that AgNO_3 is an expensive chemical, what do you think would be a better limiting reagent to use, AgNO_3 or NaCl ? Why?

Experiment Limiting Reactant Answers

P.S. Raghavan



Experiment Limiting Reactant Answers:

Visualization in Science Education John K. Gilbert, 2005-07-05 Visualization meaning both the perception of an object that is seen or touched and the mental imagery that is the product of that perception is believed to be a major strategy in all thought. It is particularly important in science which seeks causal explanations for phenomena in the world as experienced. Visualization must therefore play a major role in science education. This book addresses key issues concerning visualization in the teaching and learning of science at any level in educational systems. Visualization in Science Education draws on the insights from cognitive psychology, science and education by experts from Australia, Israel, Slovenia, UK and USA. It unites these with the practice of science education, particularly the ever increasing use of computer managed modelling packages especially in chemistry. The first section explores the significance and intellectual standing of visualization. The second section shows how the skills of visualization have been developed practically in science education. This is followed by accounts of how the educational value of visualization has been integrated into university courses in physics, genomics and geology. The fourth section documents experimental work on the classroom assessment of visualization. An endpiece summarises some of the research and development needed if the contribution of this set of universal skills is to be fully exploited at all levels and in all science subjects.

Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques. The manual helps students understand the timing and situations for the various techniques. The Beran lab manual has long been a market leading lab manual for general chemistry. Each experiment is presented with concise objectives, a comprehensive list of techniques and detailed lab intros and step by step procedures.

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O-level Chemistry Challenging Practice Solutions (Yellowreef) Thomas Bond, Chris Hughes, 2013-11-20 actual GCE exam question types must have critical resource for students and tutors. All trick question types since 2003 covered full and complete step by step solutions. Complete edition eBook available.

Guided Inquiry Experiments for General Chemistry Nancy K. Kerner, Ram S. Lamba, 2007-10-19 The use of the laboratory is a valuable tool in developing a deeper understanding of key chemical concepts from the experimental process. This lab manual encourages scientific thinking, enabling readers to conduct investigations in chemistry. It shows how to think about the processes they are investigating rather than simply performing a laboratory experiment to the specifications set by the manual. Each experiment begins with a problem scenario and ends with questions requiring feedback on the problem.

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chemistry this useful text will provide up to date experiments putting the science into context for the students **Chemistry**
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simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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Illustrated Guide to Home Chemistry Experiments Robert Bruce Thompson, 2008-04-29 For students DIY hobbyists and science buffs who can no longer get real chemistry sets this one of a kind guide explains how to set up and use a home chemistry lab with step by step instructions for conducting experiments in basic chemistry not just to make pretty colors and stinky smells but to learn how to do real lab work Purify alcohol by distillation Produce hydrogen and oxygen gas by electrolysis Smelt metallic copper from copper ore you make yourself Analyze the makeup of seawater bone and other common substances Synthesize oil of wintergreen from aspirin and rayon fiber from paper Perform forensics tests for fingerprints blood drugs and poisons and much more From the 1930s through the 1970s chemistry sets were among the most popular Christmas gifts selling in the millions But two decades ago real chemistry sets began to disappear as manufacturers and retailers became concerned about liability

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Holt Chemfile B Microscale Exp/Te 2006 Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2006

Investigating Chemistry Matthew Johll, 2008-12-22 In its new second edition Investigating Chemistry A Forensic Science Perspective remains the only

book that uses the inherently fascinating topics of crime and criminal investigations as a context for teaching the fundamental chemical concepts most often covered in an introductory nonmajors course. Covering all the standard topics, Matthew Johll capitalizes on the surge of interest in the scientific investigation of crime as sparked by CSI and other television shows, bringing together the theme of forensic science and the fundamentals of chemistry in ways that are effective and accessible for students. This edition features refined explanations of the chemical concepts which are the core of the book, as well as a more thoroughly integrated forensic theme, updated features, and an expanded media supplements package.

Ebook: Chemistry Julia Burdge, 2014-10-16. Chemistry Third Edition by Julia Burdge offers a clear writing style written with the students in mind. Julia uses her background of teaching hundreds of general chemistry students per year and creates content to offer more detailed explanation on areas where she knows they have problems. With outstanding art, a consistent problem-solving approach, interesting applications woven throughout the chapters, and a wide range of end-of-chapter problems, this is a great third edition text.

The Zinc and Iodine Book Stephen DeMeo, 2019-05-15. This book is for chemistry teachers who are thinking about reinventing their laboratory experiments that they provide to their students. More than a collection of experiments, it is an example of using a chemical theme to teach chemistry. Instead of introducing many different chemicals per experiment as is the norm in most lab manuals, this novel resource focuses on two commonly found elements, Zinc and Iodine. So what is so special about these elements? At the heart of this resource is a colorful cyclic reaction between zinc and iodine, one that produces a compound that can decompose back to its original elements. This unique phenomenon demonstrates that matter not only changes but is also conserved through a chemical reaction. Knowing that a compound can be the same but different than the reactants that formed it is to understand the essence of chemical change. Complementing this reaction, this book contains experimental activities that utilize the zinc and iodine theme to scaffold new concepts such as the properties of matter, solid and gas stoichiometry, equilibrium, kinetics, acid-base chemistry, and electrochemistry. This teacher-tested resource focuses on a set of safe substances that are appropriate for high school teachers who provide an advanced chemistry placement course and for college instructors teaching a first-year chemistry laboratory sequence.

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