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Stoichiometry and Percent Yield Practice

Answer the following questions about theoretical yield and percent yield. Make sure you show all work, units, and round for sig figs.

1. Use the following equation to answer the questions below.



If I start this reaction with 57.0 grams of nitrogen with excess hydrogen, how many grams of ammonia (NH_3) can theoretically be produced?

$$57.0 \text{ g N}_2 \times \frac{1 \text{ mol N}_2}{28.014 \text{ g N}_2} \times \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} \times \frac{17.031 \text{ g NH}_3}{1 \text{ mol NH}_3} = \boxed{69.3 \text{ g NH}_3}$$

If 60.0 grams of ammonia is actually produced, what was my percent yield of ammonia?

$$\frac{60.0}{69.3} \times 100 = \boxed{86.6\%}$$

2. Use the following equation to answer the questions below.



If 11.3 grams of carbon dioxide gas is used in this reaction, how many grams of lithium carbonate can theoretically be produced?

$$11.3 \text{ g CO}_2 \times \frac{1 \text{ mol CO}_2}{44.009 \text{ g CO}_2} \times \frac{1 \text{ mol Li}_2\text{CO}_3}{1 \text{ mol CO}_2} \times \frac{73.87 \text{ g Li}_2\text{CO}_3}{1 \text{ mol Li}_2\text{CO}_3} = \boxed{19.0 \text{ g Li}_2\text{CO}_3}$$

If 14.4 grams of lithium carbonate are actually produced, what was my percent yield?

$$\frac{14.4}{19.0} \times 100 = \boxed{75.8\%}$$

Experiment 7 Reaction Stoichiometry Percent Yield Answers

Chris Hren, Peter J. Mikulecky



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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