

Study Guide – Unit 1 Periodic Table **ANSWERS**

Structure of an atom

- List the three subatomic particles which make up an atom. Include their location, charge, and atomic mass.
 - Electron** Negative charge Electron Cloud
 - Proton** Positive charge Nucleus
 - Neutron** Neutral charge Nucleus
- In a neutral atom, which two subatomic particles are equal in number to each other?
Protons and Electrons
- What does the atomic number of an element indicate?
The number of protons in the atom of that element
- What is an isotope?
Atoms with the same number of protons, but different masses due to a different number of neutrons
- Will two different atoms of carbon always have the same number of protons?
Yes

Element key

13
Al
26.982
Aluminum

- What are the number of protons, neutrons, and electrons in an atom of aluminum?
 - Protons = 13**
 - Neutrons = 14**
 - Electrons = 13**

Bohr Model

- How do you know how many energy levels (orbits) a Bohr model will have?
 - By the period the element is located in**
- How many electrons can each energy level hold?
1st – 2 2nd – 8 3rd – 18

The Periodic Table

- How is the periodic table arranged? **Increasing atomic number**
- What are the vertical columns \updownarrow called? **Groups or Families**
- What are the horizontal rows \leftrightarrow called? **Periods**
- Why are elements put into groups? **Same number of valence electrons**
- What is a valence electron? **Number of electrons on the outer energy level (shell)**
- Know how many valence electrons each group has
- Which group of metals are highly reactive and why? **Alkali Metals. Only 1 ve⁻ to lose**
- Which group of non-metals are highly reactive and why? **Halogens. Only 1 ve⁻ to gain**
- Name some physical properties of metals: **luster, ductile, malleable, conductor of heat and electricity**
- Name some physical properties of non metals: **no luster, brittle, not a good conductor**
- What are metalloids and where are they found on the periodic table? **Share properties of metals and metalloids. On zig-zag line starting at Boron.**

Excel Lesson 3 Study Guide With Answers

Sharon Rumble



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MCTS Self-Paced Training Kit (Exam 70-667) Dan Holme, Dan Holme and Alistair Matthews, 2011 *Intermediate Microeconomics* Steve Erfle, 2017-09-14 The overarching premise of this text is that microeconomics is most effectively learned in an active learning interactive environment Students have access to more than 200 Interactive Excel Figures in the online text that allow them to move the graphs using sliders and click boxes This interactivity helps students understand how graphic elements relate to one another These files do not require knowledge of Excel More figures than are typical and many of the figures involve multiple scenarios of the same basic graph Often the text employs interactive questions that require interpreting these scenarios questions posed are answered at the bottom of the page Despite the geometric orientation this text is not light on algebraic analysis The geometry is backed up by the relevant algebra More than 500 equations are numbered for easy reference both within and across chapters And just like the geometry the algebra is essentially error free because it was used to create the graphs The geometric orientation is perfect for the non calculus enhanced classroom but the text can be readily used in a calculus based class because a calculus treatment of the material is provided in appendices and endnotes and calculus based problems are included in the *Intermediate Microeconomics An Interactive Approach Workbook*

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