

# Sulfur clock lab

## Experimental procedure

1. Measurement of the formation of visible product
2. Measurement of the initial concentration of reactants
3. Prepare solutions A, B, C, D and E as shown:

see Table 10.10.10.10

	Reactants	Reagents	Apparatus
A	1 mL	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$	Reaction setup
B	1 mL	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$	0.1 M, 0.01 M, 0.01 M
C	1 mL	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$	Reaction setup
D	1 mL	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$	Reaction setup

4. Add the solution using a graduated cylinder with completely standard chemical formulae for the reaction.
5. Transfer exactly the solution into a 100 mL beaker in the reaction mixture. Transfer the solution into a 100 mL beaker with 0.1 M  $\text{H}_2\text{O}_2$  and 0.01 M  $\text{KSCN}$  in the reaction mixture.
6. Add exactly 100 mL of the reaction mixture at the 100 mL mark of the flask.
7. Prepare and insert the flask immediately.
8. Prepare solution B, C, D, E as shown:
  - a. Transfer 100 mL of 0.1 M  $\text{H}_2\text{O}_2$  into a beaker.
  - b. Add 100 mL of 0.01 M  $\text{KSCN}$  into the beaker using a graduated cylinder.
  - c. Add the 100 mL of 0.1 M  $\text{H}_2\text{O}_2$  into the beaker.
  - d. Add the 100 mL of 0.01 M  $\text{KSCN}$  into the beaker.
9. Transfer solution A into the reaction mixture in a 100 mL beaker. Do not stir the reaction with any efficient mixer.

# Lab Sulfur Clock Answers

**Michael Brown**



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