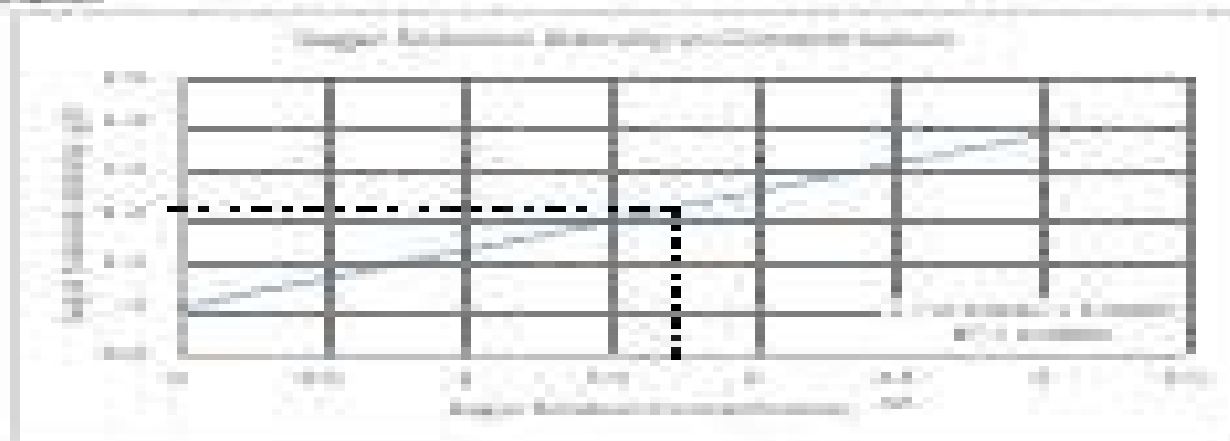


Laboratory Measurements II: Density of a Salt Solution

This experiment will build confidence in planning readings and clearly reported in the "Density and Molarity" experiment. In it, you will make a solution of salt from your graphing data, and determine experimentally density, molarity, and mass percent solution.

Substances and Their Properties

When you dissolve something like salt in water into water, the water is changing into the water-salt solution. The water molecules separate slightly from each other, and the salt molecules start pulling into it so the solution is denser than water. The solution's density goes up proportionally. For example, considering the following data for sugar solution, you can determine:



Sugar solution density values determined, the line on the graph shows the concentration between a sugar solution, concentration, density. There is a linear relationship between concentration and density. The equation for the line is $y = 0.03x + 1.00$. This graph shows the density of solution with concentration 0.2 M (0.2 moles per liter) is 1.03 g/mL.

Based on the data, which equation, you will also have a solution of sugar, which is the slope through the concentration of the solution. The line is the same. The equation for the line is $y = 0.03x + 1.00$. The line is the same. The equation for the line is $y = 0.03x + 1.00$. The line is the same. The equation for the line is $y = 0.03x + 1.00$.

* When plotting or solving for concentration, you will also have a solution of sugar, which is the slope through the concentration of the solution. The line is the same. The equation for the line is $y = 0.03x + 1.00$.

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