

Solid State

Types of solids

Crystalline solid	Amorphous solids
- Definite & regular arrangement - True solids - Long ordered arrangement of particles - These are anisotropic - Sharp M.P - Clean cleavage cut - Ex - NaCl, KCl	- Doesn't have regular arrange - Pseudo solids or super cooled liq - Short ordered arrangement - These are isotropic - Melt over certain range of temp - Irregular cleavage cut - Ex - glass, rubber

Isomorphism - Two or more substance having same crystal structure
 - same atomic ratio
 Ex NaF & MgO - 1:1
 NaNO₃ & CaCO₃ - 1:1:3

Polymorphism - A single substance that exists in two or more forms of crystalline structure
 Ex Calcite & aragonite are forms of calcium carbonate
 Ex - α-quartz, β-quartz & cristobalite - these forms of silica
 - Polymorphism occurring in elements are called allotropy
 Ex Carbon - diamond, graphite, fullerene.

Classification of crystalline solids

1) **Ionic** - Electrostatic force of attraction
 Ion - Ion force
 - Brittle, hard, high melting point
 - non conductors of electricity "when dissolved or melted in water."
 Ex - NaCl, KCl, MgCl₂

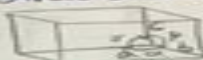
2) **Molecular** -
 - Dispersion / Dipole-Dipole / H-Bond
 - Soft, low melting point
 - non conductors & good insulators
 Ex H₂O, Br₂, CO₂, CH₄

3) **Covalent** -
 - Covalent Bond
 - Hard, high melting point, B.P
 - Most incompressible
 - Poor conductor
 Ex - C-diamond, SiO₂, BN, Carborandum

4) **Metallic** -
 - Metallic Bond
 - Variable hardness, M.P, conductivity
 - They are malleable, ductile & good conductors
 Ex - Na, Zn, Cu, Fe

Crystal Lattice - or space lattice - Lattice is a geometrical arrangement of points in 3D form such constituent particles that are attached to the lattice points are crystal lattice.

Unit Cell - The smallest repeating structural unit of a crystalline solid



→ of unit cell.

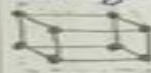
1) **Simple or primitive unit cell**

- Particles are present at corners only
 $8 \times \frac{1}{8} = 1 \text{ atom}$

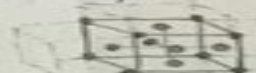
2) **Face centered** - The particles are present at the corners & also at the center
 $8 \times \frac{1}{8} + 6 \times \frac{1}{2} = 1 + 3 = 4 \text{ atoms}$

3) **Body centered** - The particles are present at the corners as well as at the centre of the unit cell
 $8 \times \frac{1}{8} + 1 = 1 + 1 = 2 \text{ atoms}$

Note -
 • An atom at the corner is shared by eight unit cell - $\frac{1}{8}$
 • An atom at the face is shared by two cell - $\frac{1}{2}$
 • An atom at the body centre of a unit cell belongs entirely to it - 1



$8 \times \frac{1}{8} = 1$
 1 atom
 SC - 1



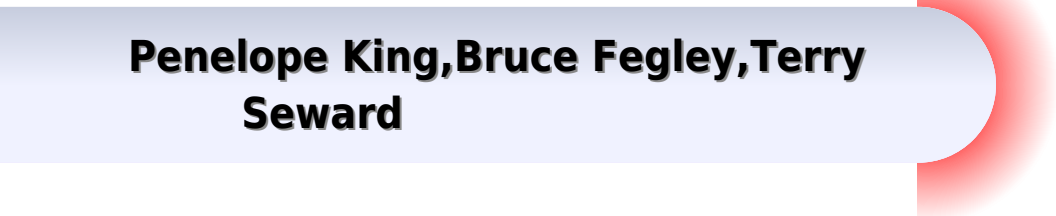
$8 \times \frac{1}{8} = 1$
 $6 \times \frac{1}{2} = 3$
 $1 + 3 = 4 \text{ atoms}$
 FCC - 4



$8 \times \frac{1}{8} = 1$
 $1 \times 1 = 1$
 $1 + 1 = 2 \text{ atoms}$
 BCC - 2

Hsc Chemistry Solid State Notes

**Penelope King, Bruce Fegley, Terry
Seward**



Hsc Chemistry Solid State Notes:

Hsc Chemistry Solid State Notes Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Hsc Chemistry Solid State Notes**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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