

Brensted-Lowry Acid/Base Theory

- Acids: Proton (H^+) donor
- Bases: Proton (H^+) acceptor
- Conjugate Acids-Base Pair: Two substances that differ only by the presence of an H^+
 - Conjugate Acid: Formed when a proton is added to a base (positive charge)
 - Add a Hydrogen to the formula and add +1 charge
 - Conjugate Base: Formed when a proton is removed from an acid (negative charge)
 - Take away a Hydrogen from the formula and subtract -1 charge
- Amphoteric: The ability for a compound to act as an acid OR as a base (EX: H_2O)
- Acids are classified based on how many H^+ are present
 - Monoprotic: 1 H^+
 - Diprotic: 2 H^+
 - Triprotic: 3 H^+
 - Diprotic and triprotic acids are also amphoteric because they have more than one proton that they can lose

Properties of Strong and Weak Acids/Bases

Strong Acids/Bases	Weak Acids/Bases
• Strong electrolytes • Ionize completely in water • Equilibrium is not present ◦ Very unlikely that the ions reform into products • Very weak attraction between ions ◦ Split apart easily	• Weak electrolytes • Does not ionize completely in water • Equilibrium exists • Strong attraction between ions ◦ Do not split apart easily
Strong acid 	Weak acid 

Equilibrium And Acid Base Study Guide

Ying-Ying Zheng



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