

2. Magnesium carbonate reacts with hydrochloric acid.



- (a) What mass of magnesium carbonate would react with excess hydrochloric acid to produce 240cm^3 of carbon dioxide, measured at room temperature and pressure?

Data: 1 mol of any gas occupies 24.0dm^3 at room temperature and pressure

Molar mass of magnesium carbonate = 84.3g mol^{-1}

- ☒ A 0.843 g
☐ B 8.43 g
☐ C 84.3 g
☐ D 843 g

$$\begin{aligned} \text{Volume of CO}_2 &= 240\text{cm}^3 = 0.24\text{dm}^3 \\ \text{moles of CO}_2 &= \frac{\text{Volume in dm}^3}{24} = \frac{0.24}{24} = 0.01 \\ \text{MgCO}_3 &\text{ contains same moles as CO}_2 \\ \text{moles} &= \frac{\text{mass}}{\text{Mr}} ; \text{ mass} = \text{moles} \times \text{Mr} \\ &= 0.01 \times 84.3 \\ &= 0.843\text{g} \end{aligned}$$

- (b) What is the minimum mass of magnesium carbonate needed to neutralise 50.0cm^3 of 0.250mol dm^{-3} hydrochloric acid?

- ☐ A 0.423 g
☒ B 0.527 g
☐ C 1.05 g
☐ D 2.11 g

$$\begin{aligned} \text{Concentration of HCl} &= \frac{\text{moles}}{\text{Vol in dm}^3} \\ \text{moles} &= \text{Conc} \times \text{Vol in dm}^3 \\ &= 0.250 \times 0.05 \\ &= 0.0125\text{moles of HCl} \end{aligned}$$

$$\begin{aligned} \text{MgCO}_3 : \text{HCl} \\ 1 : 2 \\ \frac{1}{2} : 1 \\ \frac{1}{2} \times 0.0125 : 0.0125 \\ 6.25 \times 10^{-3} : 0.0125 \\ \downarrow \qquad \qquad \downarrow \\ \text{MgCO}_3 \qquad \text{HCl} \\ \text{moles} \qquad \text{moles} \end{aligned}$$

- (c) What would be seen at the end of the reaction with excess acid?

- ☒ A A colourless solution
☐ B A coloured solution
☐ C A white precipitate
☐ D A coloured precipitate

all MgCO_3 is utilized
and color of
 MgCO_2 colorless

$$\text{mass} = \text{moles} \times \text{Mr}$$

(Total for Question 2 = 3 marks)

$$\begin{aligned} &= 6.25 \times 10^{-3} \times 84.3 \\ &= 0.527\text{g} \end{aligned}$$

Multiple Choice Questions On Stoichiometry

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Table of Contents Multiple Choice Questions On Stoichiometry

1. Understanding the eBook Multiple Choice Questions On Stoichiometry
 - The Rise of Digital Reading Multiple Choice Questions On Stoichiometry
 - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Questions On Stoichiometry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Multiple Choice Questions On Stoichiometry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Questions On Stoichiometry
 - Personalized Recommendations
 - Multiple Choice Questions On Stoichiometry User Reviews and Ratings
 - Multiple Choice Questions On Stoichiometry and Bestseller Lists
5. Accessing Multiple Choice Questions On Stoichiometry Free and Paid eBooks
 - Multiple Choice Questions On Stoichiometry Public Domain eBooks
 - Multiple Choice Questions On Stoichiometry eBook Subscription Services
 - Multiple Choice Questions On Stoichiometry Budget-Friendly Options
6. Navigating Multiple Choice Questions On Stoichiometry eBook Formats

- ePub, PDF, MOBI, and More
- Multiple Choice Questions On Stoichiometry Compatibility with Devices
- Multiple Choice Questions On Stoichiometry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Multiple Choice Questions On Stoichiometry
 - Highlighting and Note-Taking Multiple Choice Questions On Stoichiometry
 - Interactive Elements Multiple Choice Questions On Stoichiometry
- 8. Staying Engaged with Multiple Choice Questions On Stoichiometry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Multiple Choice Questions On Stoichiometry
- 9. Balancing eBooks and Physical Books Multiple Choice Questions On Stoichiometry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Multiple Choice Questions On Stoichiometry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Multiple Choice Questions On Stoichiometry
 - Setting Reading Goals Multiple Choice Questions On Stoichiometry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Multiple Choice Questions On Stoichiometry
 - Fact-Checking eBook Content of Multiple Choice Questions On Stoichiometry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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