

Gas Stoichiometry: Multiple Choice

1. What volume of pure oxygen gas, O_2 , measured at 546 K and 101.325 kPa is formed by the complete decomposition of 0.200 mol Ag_2O according to the following equation:



- A) 1.12 L B) 4.48 L C) 2.24 L D) 8.96 L E) 17.82 L

ANS: B

1. How many moles of $KClO_3$ are needed to form 2.8 L of O_2 , measured at STP, according to the following reaction: $2KClO_3 \rightarrow 2KCl + 3O_2$

- A) 0.083 mol B) 0.25 mol C) 0.17 mol D) 0.50 mol E) 0.33 mol

ANS: A

1. What volume of chlorine gas, measured at 0°C and 101.325 kPa, is required to react completely with 5.61 g of KOH in the following reaction: $Cl_{2(g)} + 2KOH \rightarrow KCl + KClO + H_{2O}$

- A) 0.56 L B) 4.48 L C) 1.12 L D) 22.4 L E) 2.24 L

ANS: C

1. Sodium hydride reacts with excess water to produce sodium hydroxide (aq) and hydrogen gas: $NaH(s) + H_2O(l) \rightarrow NaOH(aq) + H_2(g)$

A sample of NaH weighing _____ g will produce 982 mL of gas at 28.0 °C and 765 torr, when the hydrogen is collected over water. The vapor pressure of water at this temperature is 28 torr.

- A) 2.93 B) 0.960 C) 0.925 D) 0.0388 E) 925

ANS: C

1. At STP, how many liters of oxygen are required to react completely with 3.6 liters of hydrogen to form water? $2H_2(g) + O_2(g) \rightarrow 2H_2O(g)$

- A) 1.8 L B) 3.6 L C) 2.0 L D) 2.4 L

ANS: A

1. The molar volume of a gas at STP is _____ L.

- A) 0.08206 B) 62.36 C) 1.00 D) 22.4 E) 14.7

ANS: D

1. If 0.5 L of $O_2(g)$ reacts with H_2 to produce 1 L of $H_2O(g)$, what is the volume of $H_2O(g)$ obtained from 1 L of $O_{2(g)}$?

- A) 0.5 L B) 1.5 L C) 2 L D) 2.5 L

ANS: C

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Table of Contents Multiple Choice Questions On Chemical Stoichiometry

1. Understanding the eBook Multiple Choice Questions On Chemical Stoichiometry
 - The Rise of Digital Reading Multiple Choice Questions On Chemical Stoichiometry
 - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Questions On Chemical Stoichiometry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Multiple Choice Questions On Chemical Stoichiometry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Questions On Chemical Stoichiometry

- Personalized Recommendations
- Multiple Choice Questions On Chemical Stoichiometry User Reviews and Ratings
- Multiple Choice Questions On Chemical Stoichiometry and Bestseller Lists
- 5. Accessing Multiple Choice Questions On Chemical Stoichiometry Free and Paid eBooks
 - Multiple Choice Questions On Chemical Stoichiometry Public Domain eBooks
 - Multiple Choice Questions On Chemical Stoichiometry eBook Subscription Services
 - Multiple Choice Questions On Chemical Stoichiometry Budget-Friendly Options
- 6. Navigating Multiple Choice Questions On Chemical Stoichiometry eBook Formats
 - ePub, PDF, MOBI, and More
 - Multiple Choice Questions On Chemical Stoichiometry Compatibility with Devices
 - Multiple Choice Questions On Chemical Stoichiometry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Multiple Choice Questions On Chemical Stoichiometry
 - Highlighting and Note-Taking Multiple Choice Questions On Chemical Stoichiometry
 - Interactive Elements Multiple Choice Questions On Chemical Stoichiometry
- 8. Staying Engaged with Multiple Choice Questions On Chemical Stoichiometry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Multiple Choice Questions On Chemical Stoichiometry
- 9. Balancing eBooks and Physical Books Multiple Choice Questions On Chemical Stoichiometry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Multiple Choice Questions On Chemical Stoichiometry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Multiple Choice Questions On Chemical Stoichiometry
 - Setting Reading Goals Multiple Choice Questions On Chemical Stoichiometry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Multiple Choice Questions On Chemical Stoichiometry

- Fact-Checking eBook Content of Multiple Choice Questions On Chemical Stoichiometry
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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