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# Isimple Gateway Manual Toyota

**Chao Zhang**



**Isimple Gateway Manual Toyota:**

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web 1 consider the reaction  $\text{i}_2\text{O}_5 + 5 \text{CO} \rightarrow 5 \text{CO}_2 + \text{I}_2$  g a 80.0 grams of iodine pentoxide  $\text{i}_2\text{O}_5$  reacts with 28.0 grams of carbon monoxide  $\text{CO}$  determine the mass of iodine  $\text{I}_2$  which could be produced b if in the above situation only 0.160 moles of iodine  $\text{I}_2$  was produced i what mass of iodine was produced

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web limiting reactant theoretical and percentage yields key  $4 \text{KO}_2 + 2 \text{H}_2\text{O} \rightarrow 4 \text{KOH} + 3 \text{O}_2$  g if a reaction vessel contains 0 mol  $\text{KO}_2$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage of the theoretical yield

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  

$$8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$$
  

$$8.3 \text{ mole } \text{H}_2\text{O} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$$
  
 149.4 g  $\text{H}_2\text{O}$  is your theoretical yield so percent yield  

$$\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$$
  
 problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$
  
 if the reaction of 21.3 grams of  $\text{O}_2$  produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  

$$21.3 \text{ g } \text{O}_2 \times \frac{1 \text{ mole } \text{O}_2}{32 \text{ g } \text{O}_2} = 0.666 \text{ moles } \text{O}_2$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained  

$$\text{calculated mass of the product} \times 100$$
  
 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction  

$$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$$
  
 b determine the theoretical yield of  $\text{Al}_2\text{O}_3$   

$$160.0 \text{ g } \text{Al} \times \frac{1 \text{ mole } \text{Al}}{27 \text{ g } \text{Al}} \times \frac{2 \text{ moles } \text{Al}_2\text{O}_3}{4 \text{ moles } \text{Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 153.6 \text{ g } \text{Al}_2\text{O}_3$$
  
 c determine the percent yield  

$$\frac{126.0 \text{ g } \text{Al}_2\text{O}_3}{153.6 \text{ g } \text{Al}_2\text{O}_3} \times 100 = 82.0\%$$
  
 theoretical 153.6 g

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80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$  xs 1 333 8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g co 1 mol co

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