

SECTION - I

Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question :

[7]

- (i) The Arrhenius equation is _____.
- (a) $k = A e^{\frac{RT}{E_a}}$ (b) $A = k e^{\frac{-E_a}{RT}}$
- (c) $k = A e^{\frac{-RT}{E_a}}$ (d) $k = A e^{\frac{-K_a}{RT}}$
- (ii) If the enthalpy of vaporisation of water at 100°C is 186.5 J . mol⁻¹, the entropy of vaporization will be _____.
- (a) 4.0 J . K⁻¹. mol⁻¹ (b) 3.0 J . K⁻¹. mol⁻¹
- (c) 1.5 J . K⁻¹. mol⁻¹ (d) 0.5 J . K⁻¹. mol⁻¹
- (iii) The atomicity of sulphur in orthorhombic sulphur is _____.
- (a) 8 (b) 6 (c) 4 (d) 2
- (iv) The major binding force in diamond is _____.
- (a) Covalent bond (b) Ionic bond
- (c) Metallic bond (d) Co-ordinate covalent bond
- (v) The boiling point of water at high altitude is low, because _____.
- (a) the temperature is low. (b) the atmospheric pressure is low.
- (c) the temperature is high. (d) the atmospheric pressure is high.
- (vi) The molar conductivity of cation and anion of salt BA are 180 and 220 mhos respectively. The molar conductivity of salt BA at infinite dilution is _____.
- (a) 90 mhos.cm². (b) 110 mhos.cm².mol⁻¹
- (c) 200 mhos.cm².mol⁻¹ (d) 400 mhos.cm².mol⁻¹
- (vii) What is the process in which concentrated ore is reduced to the corresponding metal by heating at high temperature with a reducing agent?
- (a) Polling (b) Pyrometallurgy (c) Hydrometallurgy (d) Calcination

Q. 2. Answer any THREE of the following :

[9]

- (i) Describe anomalous behaviour of oxygen as compared with other elements of group 16 with reference to –
- (a) Magnetic property (b) Oxidation state (c) Hydrides
- (ii) What is the value of $\Delta S_{\text{surroundings}}$ for the following reaction at 298 K –
- $$6\text{CO}_{2(g)} + 6\text{H}_2\text{O}_{(l)} \longrightarrow \text{C}_6\text{H}_{12}\text{O}_{6(s)} + 6\text{O}_{2(g)}$$
- Given that : $\Delta G^\circ = 2879 \text{ kJ mol}^{-1}$, $\Delta S = -210 \text{ JK}^{-1} \text{ mol}^{-1}$.
- (iii) Sucrose decomposes in acid solution to give glucose and fructose according to the first order rate law. The half life of the reaction is 3 hours. Calculate fraction of sucrose which will remain after 8 hours.
- (iv) A solution containing 0.73 g of camphor (molar mass 152 g . mol⁻¹) in 36.8 g of acetone (boiling point 56.3°C) boils at 56.55°C. A solution of 0.564 g of unknown compound in the same weight of acetone boils at 56.46°C. Calculate the molar mass of the unknown compound.

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1. Understanding the eBook Hscmaharashtra Boardquestion Paper Oct2014 Botany
 - The Rise of Digital Reading Hscmaharashtra Boardquestion Paper Oct2014 Botany
 - Advantages of eBooks Over Traditional Books
2. Identifying Hscmaharashtra Boardquestion Paper Oct2014 Botany
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hscmaharashtra Boardquestion Paper Oct2014 Botany
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hscmaharashtra Boardquestion Paper Oct2014 Botany
 - Personalized Recommendations
 - Hscmaharashtra Boardquestion Paper Oct2014 Botany User Reviews and Ratings
 - Hscmaharashtra Boardquestion Paper Oct2014 Botany and Bestseller Lists
5. Accessing Hscmaharashtra Boardquestion Paper Oct2014 Botany Free and Paid eBooks
 - Hscmaharashtra Boardquestion Paper Oct2014 Botany Public Domain eBooks
 - Hscmaharashtra Boardquestion Paper Oct2014 Botany eBook Subscription Services
 - Hscmaharashtra Boardquestion Paper Oct2014 Botany Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Hscmaharashtra Boardquestion Paper Oct2014 Botany Compatibility with Devices
 - Hscmaharashtra Boardquestion Paper Oct2014 Botany Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hscmaharashtra Boardquestion Paper Oct2014 Botany
 - Highlighting and Note-Taking Hscmaharashtra Boardquestion Paper Oct2014 Botany
 - Interactive Elements Hscmaharashtra Boardquestion Paper Oct2014 Botany
8. Staying Engaged with Hscmaharashtra Boardquestion Paper Oct2014 Botany
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hscmaharashtra Boardquestion Paper Oct2014 Botany
9. Balancing eBooks and Physical Books Hscmaharashtra Boardquestion Paper Oct2014 Botany
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hscmaharashtra Boardquestion Paper Oct2014 Botany
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Hscmaharashtra Boardquestion Paper Oct2014 Botany
 - Setting Reading Goals Hscmaharashtra Boardquestion Paper Oct2014 Botany
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Hscmaharashtra Boardquestion Paper Oct2014 Botany
 - Fact-Checking eBook Content of Hscmaharashtra Boardquestion Paper Oct2014 Botany
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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